

CHEMISTRY MULTIPLE CHOICE QUESTIONS

A. Atoms, molecules and stoichiometry

2002 -2019



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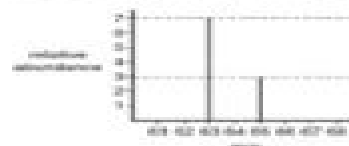
As a simplification, an adult human can be considered to have a daily diet of 1.80 kg of carbohydrate (empirical formula $C_6H_{12}O_6$).

Which mass of carbon dioxide does a person produce each day if all the carbohydrate eaten is digested and oxidised?

- A. 0.907 kg B. 0.900 kg C. 1.30 kg D. 2.54 kg

(2002 NA/12)

The diagram shows the mass spectrum of a sample of naturally-occurring copper.



What is the relative atomic mass of this copper?

- A. 63.2 B. 63.5 C. 63.6 D. 64.0

(2002 NA/10)

Which pairs of compounds have the same empirical formula?

- ethane and ethene
- ethane and cyclohexane
- cyclohexane and oct-1-ene

(2002 MC/31)

A mixture of 1.0 cm³ of methane and 1.0 cm³ of ethane was sparked with an excess of oxygen. After cooling to room temperature, the residual gas was passed through aqueous potassium hydroxide.

Which compounds have the empirical formula CH_4O ?

- methanol
- ethanoic acid
- methyl methanolate

(2002 ON/31)

What is the number of molecules in 500 cm³ of oxygen under room conditions?

- A. 1.25×10^{23} B. 1.25×10^{24} C. 3.0×10^{23} D. 3.0×10^{24}

(2008 MC/12)

In the preparation of soft margarine, glycerol trioleate reacts



is fully hydrogenated so that, on average, one of its side-chains is converted into the $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CO}_2$ moiety and two side-chains are converted into the $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CO}_2$ moiety.

How many moles of hydrogen are required to convert one mole of glycerol trioleate into the soft margarine?

- A. 4 B. 5 C. 6 D. 8

(2008 MC/22)

Use of the Data Booklet is relevant to this question.

Online classes are held every week on Wednesdays from 7.15 to 8.15 PM.

Higher Chemistry Multiple Choice Solutions 201

Julia Schneider



Higher Chemistry Multiple Choice Solutions 201:

This book delves into Higher Chemistry Multiple Choice Solutions 201. Higher Chemistry Multiple Choice Solutions 201 is a crucial topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Higher Chemistry Multiple Choice Solutions 201, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Higher Chemistry Multiple Choice Solutions 201
- Chapter 2: Essential Elements of Higher Chemistry Multiple Choice Solutions 201
- Chapter 3: Higher Chemistry Multiple Choice Solutions 201 in Everyday Life
- Chapter 4: Higher Chemistry Multiple Choice Solutions 201 in Specific Contexts
- Chapter 5: Conclusion

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6. In chapter 5, the author will draw a conclusion about Higher Chemistry Multiple Choice Solutions 201. The final chapter will summarize the key points that have been discussed throughout the book.
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