

CHEMISTRY I EQUATION SHEET

Solubility Rules

Solubility Rules			Form a Compound which is
Negative Ion Plus Positive Ion			
Any neg. ion + group 1a ions		Soluble	
Any neg. ion + Ammonium ion		Soluble	
Nitrate + any positive ion		Soluble	
Acetate	+	any positive ion except Ag ⁺ or Hg ²⁺	Soluble
bromide or iodide + Any other positive ion	+	Ag ⁺ , Pb ²⁺ , Hg ²⁺ or Cu ⁺	Not Soluble
Sulfate	+	Ca ²⁺ , Sr ²⁺ , Ba ²⁺ , Ra ²⁺ , Ag ⁺ or Pb ²⁺	Not Soluble
	+	Any other positive ion	Soluble
Sulfide	+	group 1a ions or ammonium	Soluble
	+	Be ²⁺ , Mg ²⁺ , Ca ²⁺ , Sr ²⁺ , Ba ²⁺ or Ra ²⁺	Soluble
	+	Any other positive ion	Not Soluble
Hydroxide	+	group 1a ions or ammonium	Soluble
	+	Any other positive ion	Not Soluble
Carbonate, Chromate, Phosphate, or Sulfite + Any other positive ion	+	group 1a ions or ammonium	Soluble
	+	Any other positive ion	Not Soluble

Equations

D = m/v	c = specific heat
Molarity = $\frac{\text{mol}}{\text{V}}$	d = density
$P_1 V_1 = P_2 V_2$	I = initial amount
$\frac{T_1}{T_2}$	F = final amount
Q = mc ΔT	m = mass
nRTPV =	M = molarity
F = I (1/2) *	n = number of moles & number of half lives
n = Total Time	P = pressure
Half Life Time	R = ideal gas constant
	Q = heat energy
	T = temperature
	V = volume
	X = amount of matter left

Constants

	1 atm = 101.3 kPa Pressure
	1 atm = 760 mm Hg
	R = 8.314 L x kPa / mol°C Universal
Constant or 0.0821 L x atm / mol°C	
Avogadro's number	6.02 X
Specific Heats (J / g°C)	
	Water = 4.180
	Aluminum = 0.895
	Copper = 0.387
	Iron = 0.448

Activity Series in Decreasing Activity

Metals
Lithium
Rubidium
Potassium
Calcium
Sodium
Magnesium
Aluminum
Manganese
Zinc
Iron
Nickel
Tin
Lead
Hydrogen
Copper
Silver
Platinum
Gold
Nonmetals
Fluorine
Chlorine
Bromine
Iodine

Polyatomic Ions

Name	Formula
Ammonium	NH ₄ ⁺
Acetate	C ₂ H ₃ O ₂ ⁻
Carbonate	CO ₃ ⁻²
Bicarbonate or Hydrogen carbonate	HCO ₃ ⁻
Hypochlorite	ClO ⁻
Chlorite	ClO ₂ ⁻
Chlorate	ClO ₃ ⁻
Perchlorate	ClO ₄ ⁻
Hydroxide	OH ⁻
Cyanide	CN ⁻
Nitrate	NO ₃ ⁻
Nitrite	NO ₂ ⁻
Chromate	CrO ₄ ⁻²
Oxalate	C ₂ O ₄ ⁻²
Sulfate	SO ₄ ⁻²
Sulfite	SO ₃ ⁻²
Hydrogen sulfate	HSO ₄ ⁻
Phosphate	PO ₄ ⁻³
Phosphite	PO ₃ ⁻³
Hydrogen phosphate	HPO ₄ ⁻²

Metric Conversions

Unit	Abbr.	Meaning
Mega-	M	10 ⁶
kilo-	k	10 ³
hecto-	h	10 ²
deca-	da	10 ¹
Liter	L	1
gram	g	1
meter	m	1
deci-	d	10 ⁻¹
centi-	c	10 ⁻²
milli-	m	10 ⁻³
micro-	μ	10 ⁻⁶
nano-	n	10 ⁻⁹

Calorie Conversions

Water = 1 cal/ g°C
1000 cal = 1 Cal

High School Chemistry Formulas Cheat Sheet

R Bogdan



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