

Hewlett-Packard Financial Performance Assessment

Exercise 10.17 has three similar following problems. Use all appropriate and proper scientific units. Use scientific notation for answers (e.g., 1.23×10^4).

1. Find the standard enthalpy change for the reaction $\text{C}_{10}\text{H}_{18}\text{O} + 15\text{O}_2 \rightarrow 10\text{CO}_2 + 9\text{H}_2\text{O}$ at 25°C.
- Given that $\Delta H_f^\circ(\text{C}_{10}\text{H}_{18}\text{O}) = -24.7 \text{ kJ/mol}$ and $\Delta H_f^\circ(\text{C}_{10}\text{H}_{18}) = -123.1 \text{ kJ/mol}$.

a. $\Delta H_{\text{rxn}}^\circ = \Delta H_f^\circ(\text{C}_{10}\text{H}_{18}\text{O}) + 15\Delta H_f^\circ(\text{O}_2) - 10\Delta H_f^\circ(\text{CO}_2) - 9\Delta H_f^\circ(\text{H}_2\text{O})$ $\Delta H_{\text{rxn}}^\circ = -123.1 \text{ kJ}$

b. $\Delta H_{\text{rxn}}^\circ = 10\Delta H_f^\circ(\text{CO}_2) + 9\Delta H_f^\circ(\text{H}_2\text{O}) - \Delta H_f^\circ(\text{C}_{10}\text{H}_{18}\text{O}) - 15\Delta H_f^\circ(\text{O}_2)$ $\Delta H_{\text{rxn}}^\circ = -123.1 \text{ kJ}$

2. The standard enthalpy change for the formation of aluminum oxide and aluminum hydroxide are $\Delta H_f^\circ(\text{Al}_2\text{O}_3) = -1675.7 \text{ kJ/mol}$ and $\Delta H_f^\circ(\text{Al}(\text{OH})_3) = -1201.2 \text{ kJ/mol}$. $\Delta H_f^\circ(\text{H}_2\text{O}) = -285.8 \text{ kJ/mol}$. $\Delta H_f^\circ(\text{Al}^{3+}) = -5400 \text{ kJ/mol}$. $\Delta H_f^\circ(\text{OH}^-) = -230 \text{ kJ/mol}$. Calculate the standard enthalpy change for the reaction $\text{Al}(\text{OH})_3 + 3\text{H}^+ \rightarrow \text{Al}^{3+} + 3\text{H}_2\text{O}$.

a. $\Delta H_{\text{rxn}}^\circ = 3\Delta H_f^\circ(\text{H}^+) + \Delta H_f^\circ(\text{Al}(\text{OH})_3) - \Delta H_f^\circ(\text{Al}^{3+}) - 3\Delta H_f^\circ(\text{H}_2\text{O})$ $\Delta H_{\text{rxn}}^\circ = -1201.2 \text{ kJ}$

b. $\Delta H_{\text{rxn}}^\circ = 3\Delta H_f^\circ(\text{H}_2\text{O}) + \Delta H_f^\circ(\text{Al}^{3+}) - \Delta H_f^\circ(\text{Al}(\text{OH})_3) - 3\Delta H_f^\circ(\text{H}^+)$ $\Delta H_{\text{rxn}}^\circ = -1201.2 \text{ kJ}$

c. $\Delta H_{\text{rxn}}^\circ = 3\Delta H_f^\circ(\text{H}^+) + \Delta H_f^\circ(\text{Al}(\text{OH})_3) - \Delta H_f^\circ(\text{Al}^{3+}) - 3\Delta H_f^\circ(\text{H}_2\text{O})$ $\Delta H_{\text{rxn}}^\circ = -1201.2 \text{ kJ}$

3. The combustion enthalpies for the combustion of carbon monoxide, carbon dioxide, and hydrogen, listed from top to bottom are $\Delta H_f^\circ(\text{CO}) = -110.5 \text{ kJ/mol}$, $\Delta H_f^\circ(\text{CO}_2) = -393.5 \text{ kJ/mol}$, and $\Delta H_f^\circ(\text{H}_2\text{O}) = -285.8 \text{ kJ/mol}$.

Calculate the standard enthalpy change for the reaction from the following chemical equations:

$\text{CO} + \text{H}_2\text{O} \rightarrow \text{CO}_2 + \text{H}_2$

a. $\Delta H_{\text{rxn}}^\circ = \Delta H_f^\circ(\text{CO}) + \Delta H_f^\circ(\text{H}_2\text{O}) - \Delta H_f^\circ(\text{CO}_2) - \Delta H_f^\circ(\text{H}_2)$ $\Delta H_{\text{rxn}}^\circ = -110.5 \text{ kJ}$

b. $\Delta H_{\text{rxn}}^\circ = \Delta H_f^\circ(\text{CO}_2) + \Delta H_f^\circ(\text{H}_2) - \Delta H_f^\circ(\text{CO}) - \Delta H_f^\circ(\text{H}_2\text{O})$ $\Delta H_{\text{rxn}}^\circ = -110.5 \text{ kJ}$

c. $\Delta H_{\text{rxn}}^\circ = \Delta H_f^\circ(\text{CO}_2) + \Delta H_f^\circ(\text{H}_2) - \Delta H_f^\circ(\text{CO}) - \Delta H_f^\circ(\text{H}_2\text{O})$ $\Delta H_{\text{rxn}}^\circ = -110.5 \text{ kJ}$

d. $\Delta H_{\text{rxn}}^\circ = \Delta H_f^\circ(\text{CO}) + \Delta H_f^\circ(\text{H}_2\text{O}) - \Delta H_f^\circ(\text{CO}_2) - \Delta H_f^\circ(\text{H}_2)$ $\Delta H_{\text{rxn}}^\circ = -110.5 \text{ kJ}$

Hess Law Problems With Answers

Barry McBride



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