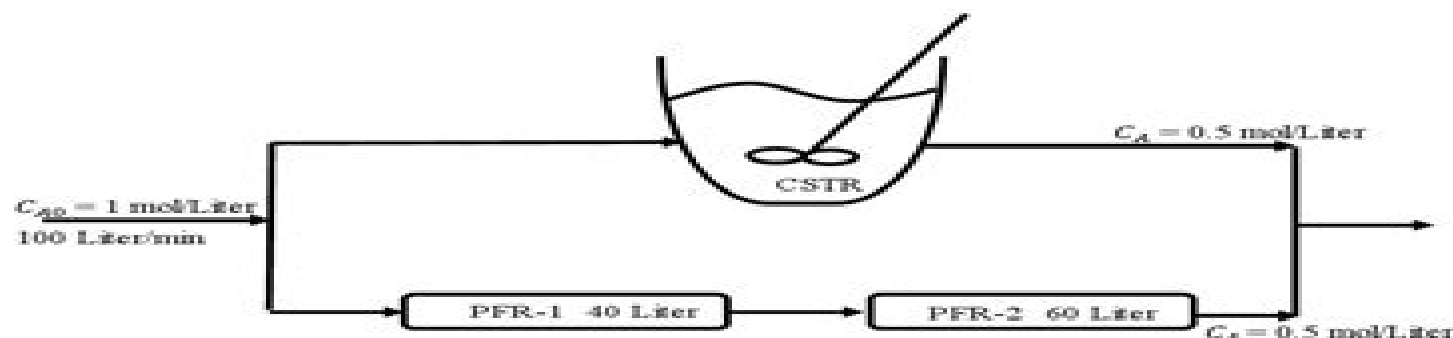


- Q.46 A first-order irreversible liquid phase reaction $A \rightarrow B$ ($k = 0.1 \text{ min}^{-1}$) is carried out under isothermal, steady state conditions in the following reactor arrangement comprising an ideal CSTR (Continuous-Stirred Tank Reactor) and two ideal PFRs (Plug Flow Reactors). From the information in the figure, the volume of the CSTR (in Liters) is _____ (rounded off to the nearest integer).



- Q.47 The elementary liquid-phase irreversible reactions



take place in an isothermal ideal CSTR (Continuous-Stirred Tank Reactor). Pure A is fed to the reactor at a concentration of 2 mol/Liter. For the residence time that maximizes the exit concentration of B, the percentage yield of B, defined as $\left(\frac{\text{net formation rate of B}}{\text{consumption rate of A}} \times 100 \right)$, is _____ (rounded off to the nearest integer).

- Q.48 The elementary irreversible gas-phase reaction $A \rightarrow B + C$ is carried out adiabatically in an ideal CSTR (Continuous-Stirred Tank Reactor) operating at 10 atm. Pure A enters the CSTR at a flow rate of 10 mol/s and a temperature of 450 K. Assume A, B and C to be ideal gases. The specific heat capacity at constant pressure ($C_{p,i}$) and heat of formation ($H_{f,i}^\circ$), of component i ($i = A, B, C$), are:

$$\begin{array}{lll} C_{pA} = 30 \text{ J/(mol K)} & C_{pB} = 10 \text{ J/(mol K)} & C_{pC} = 20 \text{ J/(mol K)} \\ H_A^\circ = -90 \text{ kJ/mol} & H_B^\circ = -54 \text{ kJ/mol} & H_C^\circ = -45 \text{ kJ/mol} \end{array}$$

The reaction rate constant k (per second) $= 0.133 \exp \left\{ \frac{E}{R} \left(\frac{1}{450} - \frac{1}{T} \right) \right\}$, where $E = 31.4$ kJ/mol and universal gas constant $R = 0.082 \text{ L atm/(mol K)} = 8.314 \text{ J/(mol K)}$. The shaft work may be neglected in the analysis, and specific heat capacities do not vary with temperature. All heats of formation are referenced to 273 K. The reactor volume (in Liters) for 75% conversion is _____ (rounded off to the nearest integer).

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Table of Contents Graduate Aptitude Test In Engineering Question Papers

1. Understanding the eBook Graduate Aptitude Test In Engineering Question Papers
 - The Rise of Digital Reading Graduate Aptitude Test In Engineering Question Papers
 - Advantages of eBooks Over Traditional Books
2. Identifying Graduate Aptitude Test In Engineering Question Papers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Graduate Aptitude Test In Engineering Question Papers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Graduate Aptitude Test In Engineering Question Papers
 - Personalized Recommendations
 - Graduate Aptitude Test In Engineering Question Papers User Reviews and Ratings
 - Graduate Aptitude Test In Engineering Question Papers and Bestseller Lists
5. Accessing Graduate Aptitude Test In Engineering Question Papers Free and Paid eBooks
 - Graduate Aptitude Test In Engineering Question Papers Public Domain eBooks
 - Graduate Aptitude Test In Engineering Question Papers eBook Subscription Services
 - Graduate Aptitude Test In Engineering Question Papers Budget-Friendly Options

6. Navigating Graduate Aptitude Test In Engineering Question Papers eBook Formats
 - ePub, PDF, MOBI, and More
 - Graduate Aptitude Test In Engineering Question Papers Compatibility with Devices
 - Graduate Aptitude Test In Engineering Question Papers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Graduate Aptitude Test In Engineering Question Papers
 - Highlighting and Note-Taking Graduate Aptitude Test In Engineering Question Papers
 - Interactive Elements Graduate Aptitude Test In Engineering Question Papers
8. Staying Engaged with Graduate Aptitude Test In Engineering Question Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Graduate Aptitude Test In Engineering Question Papers
9. Balancing eBooks and Physical Books Graduate Aptitude Test In Engineering Question Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Graduate Aptitude Test In Engineering Question Papers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Graduate Aptitude Test In Engineering Question Papers
 - Setting Reading Goals Graduate Aptitude Test In Engineering Question Papers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Graduate Aptitude Test In Engineering Question Papers
 - Fact-Checking eBook Content of Graduate Aptitude Test In Engineering Question Papers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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