

Chemistry Practice Test 1 Solution

1. Select the best word from the list to complete each of the following statements. (14 marks)

elements, gas, groups, formula, ion, neutralisation, negative, combustion,
decomposition, metals, neutrons, hydroxide, carbonate, positive

- a. CO_3^{2-} is called the carbonate ion.
- b. The formula for Sodium hydroxide is NaOH .
- c. Iron, copper and lead are classified as metals.
- d. An acid reacting with a carbonate produces bubbles of carbon dioxide gas.
- e. An atom that has gained or lost electrons is called an ion.
- f. One substance splitting into two or more substances is called decomposition.
- g. Protons and neutrons are found in the nucleus of atoms.
- h. OH^- is called the hydroxide ion.
- i. Columns in the periodic table are called groups.
- j. The reaction between an acid and a base is called neutralisation.
- k. The periodic table contains only elements.
- l. The word combustion means burning.
- m. Protons have a positive charge.
- n. Electrons have a negative charge.

2. Using the Periodic Table, complete the table below. (4 marks)

Element Name	Element Symbol	Atomic number	Number of protons	Number of electrons
Lithium	Li	3	3	3

3. Complete the following table. (6 marks)

Reaction	Reactants	Products
Precipitation	Two soluble salt solutions	A soluble salt solution and an insoluble salt
Neutralisation	Acid and base	A salt and water
Acid on Carbonate	Acid and carbonate	A salt, water and carbon dioxide gas
Acid on Metal	Acid and metal	A salt and hydrogen gas

4. Identify the mystery substance (P,Q,R,S,T,U) for each reaction. (6 marks)

Reaction	Mystery substance
Neutralisation: Hydrochloric acid is added to P and sodium chloride and water are produced.	P: Sodium hydroxide
Decomposition: Electricity is passed through water and Q and oxygen gases are produced.	Q: Hydrogen
Combustion: Wood is heated in air and a gas R and water are produced.	R: Carbon dioxide
Precipitation: Aqueous silver nitrate is added to aqueous sodium chloride and S , a white solid, is observed to have formed at the bottom of the reaction vessel.	S: Silver chloride
Acid on a carbonate: Hydrobromic acid is added to calcium carbonate and water, calcium bromide and a gas T are produced. The gas turns lime water milky.	T: Carbon dioxide
MASH: Magnesium metal is placed in acid and a gas U is formed.	U: Hydrogen

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R Bogdan

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