

**Question 1. Calculate the molecular mass of the following:**

**(i)  $\text{H}_2\text{O}$ (ii)  $\text{CO}_2$ (iii)  $\text{CH}_4$**

**Answer:** (i) Molecular mass of  $\text{H}_2\text{O} = 2(1.008 \text{ amu}) + 16.00 \text{ amu} = 18.016 \text{ amu}$

(ii) Molecular mass of  $\text{CO}_2 = 12.01 \text{ amu} + 2 \times 16.00 \text{ amu} = 44.01 \text{ amu}$

(iii) Molecular mass of  $\text{CH}_4 = 12.01 \text{ amu} + 4 (1.008 \text{ amu}) = 16.042 \text{ amu}$

**Question 2. Calculate the mass percent of different elements present in sodium sulphate ( $\text{Na}_2 \text{SO}_4$ ).**

**Answer:**

$$\text{Mass \% of an element} = \frac{\text{Mass of that element in the compound}}{\text{Molar mass of the compound}} \times 100$$

Now, 
$$\begin{aligned} \text{Molar mass of } \text{Na}_2\text{SO}_4 &= 2 (23.0) + 32.0 + 4 \times 16.0 \\ &= 142 \text{ g mol}^{-1}, \end{aligned}$$

$$\begin{aligned} \text{Mass percent of sodium} &= \frac{46}{142} \times 100 \\ &= 32.39 \% \end{aligned}$$

$$\begin{aligned} \text{Mass percent of sulphur} &= \frac{32}{142} \times 100 \\ &= 22.54 \% \end{aligned}$$

$$\begin{aligned} \text{Mass percent of oxygen} &= \frac{64}{142} \times 100 \\ &= 45.07 \% \end{aligned}$$

**Question 3. Determine the empirical formula of an oxide of Iron which has 69.9 % iron and 30.1 % dioxygen by mass.**

# Grade 11 Chemistry Lab Answers

**R Sanford**



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