

Section 7.2: Equilibrium Law and the Equilibrium Constant

Tutorial 1 Practice, page 431



Products: $2 \text{CO}(\text{g})$; $\text{O}_2(\text{g})$

Reactant: $2 \text{CO}_2(\text{g})$

$$\text{Equilibrium law equation: } K = \frac{[\text{CO}(\text{g})]^2 [\text{O}_2(\text{g})]}{[\text{CO}_2(\text{g})]^2}$$



Products: $4 \text{HCl}(\text{g})$; $\text{O}_2(\text{g})$

Reactants: $2 \text{Cl}_2(\text{g})$; $2 \text{H}_2\text{O}(\text{g})$

$$\text{Equilibrium law equation: } K = \frac{[\text{HCl}(\text{g})]^4 [\text{O}_2(\text{g})]}{[\text{Cl}_2(\text{g})]^2 [\text{H}_2\text{O}(\text{g})]^2}$$



Product: $3 \text{O}_2(\text{g})$

Reactant: $2 \text{O}_3(\text{g})$

$$\text{Equilibrium law equation: } K = \frac{[\text{O}_2(\text{g})]^3}{[\text{O}_3(\text{g})]^2}$$



Products: $2 \text{N}_2(\text{g})$; $6 \text{H}_2\text{O}(\text{g})$

Reactants: $4 \text{NH}_3(\text{g})$; $3 \text{O}_2(\text{g})$

$$\text{Equilibrium law equation: } K = \frac{[\text{N}_2(\text{g})]^2 [\text{H}_2\text{O}(\text{g})]^6}{[\text{NH}_3(\text{g})]^4 [\text{O}_2(\text{g})]^3}$$

2. **Given:** $\text{CO}(\text{g}) + 2 \text{H}_2(\text{g}) \rightleftharpoons \text{CH}_3\text{OH}(\text{g})$; $[\text{CO}(\text{g})]_{\text{equilibrium}} = 0.079 \text{ mol/L}$;

$[\text{H}_2(\text{g})]_{\text{equilibrium}} = 0.158 \text{ mol/L}$; $[\text{CH}_3\text{OH}(\text{g})]_{\text{equilibrium}} = 0.021 \text{ mol/L}$

Required: K

Solution: Write the equilibrium law equation using the balanced chemical equation. Then, substitute the equilibrium concentrations into the equilibrium law equation and solve for K .

$$\begin{aligned} K &= \frac{[\text{CH}_3\text{OH}(\text{g})]}{[\text{CO}(\text{g})][\text{H}_2(\text{g})]^2} \\ &= \frac{(0.021)}{(0.079)(0.158)^2} \end{aligned}$$

$$K = 11$$

Statement: The equilibrium constant, K , for the reaction is 11.

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