

Key

Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are *three-dimensional*. The true shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

Model 1 – Lewis Structures

Lewis Structures		H ₂ CO	3 electron domains (3 bonding, 0 nonbonding)	3-D Molecular Shape
1. H ₂ CO				<i>trigonal planar</i>
2. BeF ₂			2 electron domains (2 bonding, 0 nonbonding)	<i>Linear</i>
3. CH ₄			4 electron domains (4 bonding, 0 nonbonding)	<i>Tetrahedral</i>
4. NH ₃			4 electron domains (3 bonding, 1 nonbonding)	<i>pyramidal</i>
5. H ₂ O			4 electron domains (2 bonding, 2 nonbonding)	<i>Bent</i>
6. CO ₂			2 electron domains (2 bonding, 0 nonbonding)	<i>linear</i>

Lone pair = ..

Geometry Pogil Answer Key

VM Jensen

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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