# **Calorimetry Lab Report**

Lab Members: Dylan Craig, and Brayden

**Objective:** The objective of this lab is to find out if a fishing weight was actually made of lead, to investigate the first law of thermodynamics and Calorimetry, and to design an experiment by first creating an analytical model as a guide.

Introduction: By combining the first law of thermodynamics with Calorimetry, one can determine if a fishing weight that claims to be lead, is actually made of lead or not, The fiest law of thermodynamics states that you can change the internal energy of a system by doing work or applying heat to a system.

$$\Delta E_{em} = W_{er} + Q$$

This can help determine the hear capacity of the lead weight.

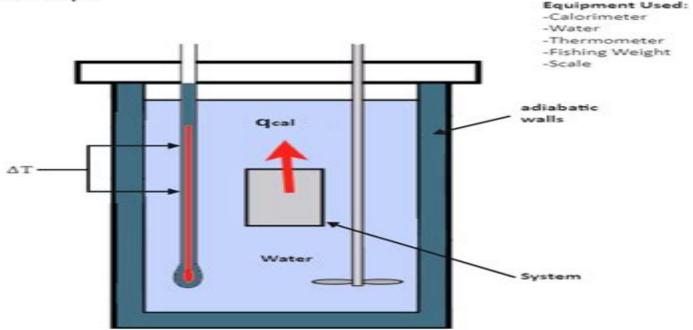
The heat [Q] is a physical quantity that characterizes the transfer of energy in the form of heat from an environment to a system (pg. 1, Calorimetry).

$$Q = mc \Delta T$$

It is important to note that when one is analyzing two objects that are exchanging heat, the conservation of energy is:

Qhottocold + Qcoldtohat=0

#### Experimental Setup:



# **Enthalpy Calorimeter Lab Report**

Carnegie Institution of Washington.
Geophysical Laboratory

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