

## Figure 15.11(a)

### Long-term equilibrium (2000) conditions

Consider your system initially before the addition of  $\text{NaOH}$ . The system is at equilibrium and the pH is 7.00.

- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.
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- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.

When you add  $\text{NaOH}$  to the  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system, the pH will increase and the ratio of  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  will decrease.

When you add  $\text{NaOH}$  to the  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system, the pH will increase and the ratio of  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  will decrease.

100

## Figure 15.11(b)

### Long-term equilibrium

- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.
- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.
- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.
- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.

### Long-term equilibrium (2000) conditions

Consider your system after the addition of  $\text{NaOH}$ . The system is at equilibrium and the pH is 7.00.

- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.
- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.
- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.
- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.

100

## Figure 15.11(c)

### Long-term equilibrium (2000) conditions

Consider your system after the addition of  $\text{NaOH}$ . The system is at equilibrium and the pH is 7.00.

When you add  $\text{NaOH}$  to the  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system, the pH will increase and the ratio of  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  will decrease.

When you add  $\text{NaOH}$  to the  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system, the pH will increase and the ratio of  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  will decrease.

## Figure 15.11(d)

### Long-term equilibrium

- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.
- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.
- The  $\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$  buffer system is at a ratio of 1:1 and the pH is 7.00.

# Motorola S9 Hd Troubleshooting

**OKSANA. CHALIFOUR**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Motorola S9 Hd Troubleshooting:**

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