

CHEMISTRY

CHEAT SHEET

CHEMICAL KINETICS

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DEFINITIONS: The study of the rates of chemical reactions is called chemical kinetics. It is concerned with the rates of chemical reactions and the factors that affect them.

Rate of Reaction: The rate of a chemical reaction is the change in concentration of a reactant or product per unit time.

Order of Reaction: The order of a chemical reaction is the sum of the powers of the concentration terms in the rate equation.

Half-life: The half-life of a reaction is the time taken for the concentration of a reactant to fall to half its initial value.

Rate Equation: The rate equation for a chemical reaction is an equation that relates the rate of reaction to the concentrations of the reactants.

Rate Constant: The rate constant is a proportionality constant that appears in the rate equation.

Activation Energy: The activation energy of a chemical reaction is the minimum energy that must be supplied to the reactants for the reaction to occur.

Reaction Orders			
Order	Description	Rate Equation	Half-life Equation
Zero order	Rate is independent of the concentration of the reactants	$\text{Rate} = k$	$t_{1/2} = \frac{[A]_0}{2k}$
First order	Rate is directly proportional to the concentration of one reactant	$\text{Rate} = k[A]$	$t_{1/2} = \frac{\ln 2}{k}$
Second order	Rate is proportional to the square of the concentration of one reactant or to the product of the concentrations of two reactants	$\text{Rate} = k[A]^2$ or $\text{Rate} = k[A][B]$	$t_{1/2} = \frac{1}{k[A]_0}$

Rate of Reaction

Rate = $\frac{\Delta[\text{Concentration}]}{\Delta \text{Time}}$

Units: $\text{mol dm}^{-3} \text{s}^{-1}$

Rate Equation

Rate = $k[A]^m[B]^n$

Units: $\text{mol dm}^{-3} \text{s}^{-1}$

Half-life

Units: s

Rate of Reaction

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Half-life

Units: s

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