

## EXPERIMENT 7

### CHEMICAL KINETICS

### SAMPLE REPORT

#### Special Notes

1. The report must be written in ink (blue pen), no pencil allowed. However, the graph can be drawn in pencil.
2. The layout of the report must be tidy (no white out) and legible in writing.  
(Work out all the solutions on some scrap papers first, then copy to the report sheets.)
3. Page 7 - 1
  - ① Record exact temperatures in °C to 1 decimal.
  - ② Both the "Time" in seconds and "Relative Rate" must be integers.
  - ③ Report the "Initial  $[I^-]$ ", "Initial  $[S_2O_8^{2-}]$ " and "Rate" in *3 s.f.*
4. Page 7 - 2
  - ① Show **all** calculations for:
    - (A) Concentration of initial  $[I^-]$ .
    - (B) Concentration of initial  $[S_2O_8^{2-}]$ .
    - (C) Rate of Reaction of  $S_2O_8^{2-}$ , and
    - (D) Relative Rates
  - ② Keep all calculations in *3 s.f.*. However, Relative Rate, Order of Reaction, and "m" and "n" must be reported as integers.
5. Page 7 - 3
  - ① For Question 1, the conclusion must include: **rate**  $\propto$   $[I^-]$ .
  - ② For Question 2, the conclusion must include: **rate**  $\propto$   $[S_2O_8^{2-}]$ .
  - ③ For Question 3, the rate law is: **rate** =  $k [I^-] [S_2O_8^{2-}]$
6. Page 7 - 4
  - ① Show calculations of rate constants,  $k$ , for all five reactions.
  - ② The units of rate constant,  $k$ , is  $M^{-1}s^{-1}$ .
  - ③ The **accuracy** of  $k$  is between  $3.30 \times 10^{-3}$  to  $3.90 \times 10^{-3} M^{-1}s^{-1}$  at 22-23°C.
  - ④ The **precision** of  $k$   $\pm 0.30 \times 10^{-3} M^{-1}s^{-1}$  between highest and lowest  $k$ .

# Kinetics Of A Chemical Reaction Lab Report

**Lauren Gardner**



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