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web limiting reactant and percent yield practice name 1 consider the following reaction $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{}_3\text{PO}_4 + \text{N}_2\text{O}$ which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

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web limiting reagents and percentage yield worksheet 1 consider the reaction $\text{I}_2\text{O}_5 + \text{CO} \rightarrow \text{I}_2 + \text{CO}_2$ 80.0 grams of iodine(V) oxide I_2O_5 reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced $80.0 \text{ g } \text{I}_2\text{O}_5 \times \frac{1 \text{ mol } \text{I}_2\text{O}_5}{1 \text{ mol } \text{I}_2} \times \frac{1 \text{ mol } \text{I}_2}{1 \text{ mol } \text{I}_2\text{O}_5} = 28.0 \text{ g } \text{I}_2$

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ label 3 7 3