

Name KEY
 Date Due _____
 Hand In With Corrections by _____

Chemistry 12
 Worksheet 4-2

82

Bronsted Acids and Equilibria

1. Write the formula for a *proton* (1 mark) H^+
2. Write the formula for a *hydrated proton* (1 mark) H_3O^+
3. Write the formula for a *hydronium* ion (1 mark) H_3O^+
4. Give the *Arrhenius* definition of an *acid* (1 mark) any subst. which releases H^+ ions in water
5. Give the *Arrhenius* definition of a *base* (1 mark) any subst. which releases OH^- ions in water.
6. Give the *Bronsted* definition of an *acid* (1 mark) a species which donates protons to another species
7. Give the *Bronsted* definition of a *base* (1 mark) a species which accepts protons from another species
8. Given the equation: $HCO_3^- + H_2S \rightleftharpoons H_2CO_3 + HS^-$
 - a) The **acid** on the left side is (1 mark) H_2S
 - b) The **base** on the left side is (1 mark) HCO_3^-
 - c) The **acid** on the right side is (1 mark) H_2CO_3
 - d) The **base** on the right side is (1 mark) HS^-

Lecture 30 Acids And Bases Solutions Worksheet

J Elliott



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