

Question 13
Acids and Bases IV
Advanced

13. Refer to the following sample of Question 12 added to a pH 7.00 saturated solution of HNO_2 . What is the concentration of HNO_2 in this solution. Assume that the total H^+ density does not change. Assume that the overall volume of the solution does not change.

$$\text{HNO}_2(aq) \rightleftharpoons \text{HNO}_3(aq) + \text{H}_2\text{O}(l)$$

$$\text{HNO}_2(aq) + \text{H}_2\text{O}(l) \rightleftharpoons \text{H}_2\text{NO}_2^+(aq)$$

$\text{HNO}_2(aq)$ with identical complexity. The addition of $\text{H}_2\text{O}(l)$ does not affect the equilibrium of the two reactions that form, according to Le Chatelier's principle. Assume $\text{H}_2\text{O}(l)$ will combine with $\text{HNO}_2(aq)$ to form the products $\text{HNO}_3(aq)$. This will continue until the product of $[\text{HNO}_2(aq)]$ and $[\text{H}_2\text{O}(l)]$ equals K_{eq} (not equal). This will decrease the concentration of HNO_2 in the solution.

13. A chemical reaction is created by combining the acid of weakly acidic. What is the pH of the solution of HNO_2 (pH = 4.2) in H_2O . Assume that the volume of the solution does not change.

- a. What is the pH of the solution?

13. Answer: $\text{pH} = 4.2$ (not 4.2). $\text{pH} = \text{pH}_0 + \log \left(\frac{[\text{HNO}_2]}{[\text{H}_2\text{O}] + [\text{HNO}_2]} \right) = -\log(4.2) + \log(4.2) = \text{pH}_0 = 4.2$

13. What is the pH of the solution of HNO_2 in solution? Assume that the volume of the solution does not change.

13. Answer: $\text{pH} = 4.2$ (not 4.2). $\text{pH} = \text{pH}_0 + \log \left(\frac{[\text{HNO}_2]}{[\text{H}_2\text{O}] + [\text{HNO}_2]} \right) = -\log(4.2) + \log(4.2) = \text{pH}_0 = 4.2$

	HNO_2	H_2O	HNO_3	H_2NO_2^+
Initial	0.01 M	0.01 M	0.01 M	0.01 M
Change	-0.01 M	-0.01 M	+0.01 M	+0.01 M
Equilibrium	0.01 M	0.01 M	0.01 M	0.01 M

13. Answer: $\text{pH} = 4.2$ (not 4.2). $\text{pH} = \text{pH}_0 + \log \left(\frac{[\text{HNO}_2]}{[\text{H}_2\text{O}] + [\text{HNO}_2]} \right) = -\log(4.2) + \log(4.2) = \text{pH}_0 = 4.2$

Lecture 32 Acids And Bases Iv Answers

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