

(e) Choose the correct answer from the four options given below:

[5]

- (i) The fusion of the sperm and ovum is termed as:
  - A. Reproduction
  - B. Development
  - C. Fertilization
  - D. Embryo
- (ii) Agranulocytes are:
  - A. Lymphocytes, Monocytes
  - B. Lymphocytes, Basophils
  - C. Eosinophils, Basophils
  - D. Eosinophils, Monocytes
- (iii) Which of the following is not a natural reflex action?
  - A. Knee-jerk
  - B. Blinking of eyes due to strong light
  - C. Salivation at the sight of food
  - D. Sneezing when any irritant enters the nose
- (iv) The structural and functional units of excretion in the human kidney is the:
  - A. Ureter
  - B. Bowman's capsule
  - C. Renal pelvis
  - D. Nephron
- (v) In a human female, ovum consists of:
  - A. 23 pairs of autosomes
  - B. 22 pairs of autosomes and 1 pair of sex chromosomes
  - C. 22 autosomes and 1 Y-chromosome
  - D. 22 autosomes and 1 X-chromosome

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temperature change plugging in the values given in the problem you get  $q_{\text{water}} = 4.18 \text{ J/g} \times 110 \text{ g} \times 26.6 \text{ }^{\circ}\text{C}$   
 $550 \text{ J}$   $\Delta h_{\text{water}} = 550 \text{ J}$

**calorimetry problems worksheet bremertonschools org** - Aug 26 2022

web calorimetry problems calorimetry problems  $q_{\text{sur}} = m \times c \times \Delta t$  heat  $m$  mass  $q_{\text{rxn}}$   $q_{\text{sur}}$   $t_f$   $t_i$  name per date  $c$  specific heat  
 for water  $4.184 \text{ J/g}^{\circ}\text{C}$  1 what is the specific heat of aluminum if the temperature of a  $28.4 \text{ g}$  sample of aluminum is increased  
 by  $8.1^{\circ}\text{C}$  when  $207 \text{ J}$  of heat is added 2

**7 3 heats of reactions and calorimetry chemistry libretexts** - Dec 30 2022

web jul 12 2023 this concept lies at the heart of all calorimetry problems and calculations because the heat released or  
 absorbed at constant pressure is equal to  $\Delta h$  the relationship between heat and  $\Delta h_{\text{rxn}}$  is  $\Delta h_{\text{rxn}} = q_{\text{rxn}} / n$   $q_{\text{calorimeter}} = mcs\Delta t$  7 3  
 14 7 3 14  $\Delta h_{\text{rxn}} = q_{\text{rxn}} / n$   $q_{\text{calorimeter}} = mcs\Delta t$

**calorimetry practice problems mrphysics org** - Mar 21 2022

web use the information in this chart to answer the following questions 1 if i have  $125 \text{ grams}$  of water at a temperature of  
 $25.0^{\circ}\text{C}$  and heat it to  $145.0^{\circ}\text{C}$  what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the  
 process in 1 3 i have  $18.7 \text{ grams}$  of an unknown liquid chemical at a temperature of  $35.0^{\circ}\text{C}$

calorimetry chem101 online general chemistry lumen - Oct 28 2022

web sample problem calorimetry and enthalpy changes in an experiment  $25.0 \text{ ml}$  of  $1.00 \text{ M HCl}$  at  $25.0^{\circ}\text{C}$  is added to  $25.0 \text{ ml}$   
 of  $1.00 \text{ M NaOH}$  at  $25.0^{\circ}\text{C}$  in a foam cup calorimeter a reaction occurs and the temperature rises to  $32.0^{\circ}\text{C}$  calculate the  
 enthalpy change in  $\text{kJ}$  for this reaction assume the densities of the solutions are  $1.00 \text{ g/ml}$  and that

6 4 calorimetry chemistry libretexts - Jan 31 2023

web before we practice calorimetry problems involving chemical reactions consider a simple example that illustrates the core  
 idea behind calorimetry suppose we initially have a high temperature substance such as a hot piece of metal  $m$  and a low  
 temperature substance such as cool water  $w$  answer the initial temperature of the copper

**calorimeter questions practice questions with answers** - Jun 23 2022

web calorimeter questions practice questions with answers explanations physics physics concept questions and answers  
 calorimeter questions calorimeter questions a calorimeter is an apparatus used for calculating the heat developed during a  
 chemical mechanical or electrical reaction it also helps to measure the heat capacity of various

**calorimetry problems thermochemistry practice specific heat** - Apr 02 2023

web dec 18 2016 this chemistry video tutorial explains how to solve calorimetry problems in thermochemistry it shows you  
 how to calculate the quantity of heat transferred

**calorimetry practice problems to post laney college** - Jul 25 2022

web calorimetry practice problems 1 when 5 000 grams of ammonia react with an excess of oxygen and  $\text{CH}_4$  in a bomb calorimeter with a total heat capacity of 15 48 kJ  $^{\circ}\text{C}^{-1}$  the temperature of the calorimeter and its contents changes from 19.77  $^{\circ}\text{C}$  to 33.90  $^{\circ}\text{C}$  calculate  $\Delta H$  for the following reaction as written  $2\text{NH}_3(\text{g}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{N}_2(\text{g}) + 6\text{H}_2\text{O}(\text{l})$

**heat capacity and calorimetry practice khan academy** - Jun 04 2023

web choose 1 answer the final temperature is closer to  $t_1$  than to  $t_2$  a the final temperature is closer to  $t_1$  than to  $t_2$  the final temperature is exactly halfway between  $t_1$  and  $t_2$  b the final temperature is exactly halfway between  $t_1$  and  $t_2$  the final temperature is closer to  $t_2$  than to  $t_1$  c

*calorimetry practice problems chemistry steps* - Sep 07 2023

web in this set of practice questions we will go over the main types of questions on calorimetry including the heat capacity the heat of reaction finding the final temperature of a mixture constant pressure calorimetry and constant volume calorimetry a few important concepts and formulas you will need to solve these questions

*calorimetry worksheet laney college* - Nov 28 2022

web calorimetry worksheet 1 if 0.315 moles of hexane  $\text{C}_6\text{H}_{14}$  is combusted in a bomb calorimeter containing 5.65 liters of water calculate the molar heat of combustion of hexane if the water temperature rises 5.4  $^{\circ}\text{C}$  the specific heat capacity of water is 4.184 J  $\text{g}^{-1} \text{ } ^{\circ}\text{C}^{-1}$   $\Delta H_{\text{comb}} =$  5 650 J  $\text{g}^{-1}$  4.184 J  $\text{g}^{-1} \text{ } ^{\circ}\text{C}^{-1}$  1310 kJ

*calorimetry problems 1* - Feb 17 2022

web 6 a 218 g sample of steam at 121  $^{\circ}\text{C}$  is cooled to ice at 14  $^{\circ}\text{C}$  find the change in heat content of the system 7 in going from ice at 34  $^{\circ}\text{C}$  to steam at 138  $^{\circ}\text{C}$  a sample of water absorbs  $1.41 \times 10^5 \text{ J}$  find the mass of the sample answers 1 2 68  $\times 10^4 \text{ J}$  2 6 23  $\times 10^4 \text{ J}$  3 1 71  $\times 10^4 \text{ J}$  4 1 11  $\times 10^5 \text{ J}$  5 2 11  $\times 10^5 \text{ J}$  6

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