

The Hydrogen Atom – Student Guide

Background Material

Carefully read the background pages created through Eureka, Slides, and Animations and answer the following questions in-depth and understanding.

Question 1: Complete the following table which compares how the Bohr model and the Quantum model compare the Hydrogen atom. In each case, they both describe things in a unique way and in some ways they do not.

Concept Name	Bohr Model	Quantum Model
Electrons	positively charged particles in the atom	positively charged particles in the atom
Orbitals	negatively charged and orbit the nucleus	negatively charged clouds
Location of the electron in the atom without	is in two fixed paths	electron cloud pattern
Energy of the electron in the atom without	electron has fixed energy levels	energy has probability
Energy of the electron in the atom without	$E = -13.6 \text{ eV}$	$E = -13.6 \text{ eV}$

Question 2: Complete the following table which shows the comparison of two different theories by putting predictions made by each theory side-by-side. The table can be completed by using the Bohr model or Quantum model. Significant differences, similar responses, similar energy, and the same results through calculations are all fine.

Hydrogen Atom Student Guide Naap Answers

**Rudolph M. Langer, California Institute
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Mathematics and Astronomy**



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Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section s critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text The Hydrogen Atom Energy Levels ,1976 *Student's Guide to Chemistry; a Modern Introduction* David Brooks,1974 The Reaction of Hydrogen Atom with Hydrogen Molecular Ion Dion Allan Stams,1985 **On the Energy Levels of a Model of the Compressed Hydrogen Atom** S. R. de Groot,C. A. ten Seldam,1946 Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e Kenneth W. Raymond,2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry Student Solutions Manual for General Chemistry John McMurry,Robert Fay,2013-03-07 By Joseph Topich Virginia Commonwealth University This manual for students contains solutions to selected all in chapter problems and even numbered end of chapter problems *Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry* Katherine J. Denniston,Denniston Katherine,2006 **The Hydrogen Atom** Rudolph M. Langer,California Institute of Technology. Division of Physics, Mathematics and Astronomy,1927 Chemistry Steven S. Zumdahl,Thomas Hummel,2015-06 This manual provides detailed solutions for half of the end of chapter exercises designated by blue question numbers using the strategies emphasized in the text This manual has been thoroughly checked for precision and accuracy Answers to the For Review questions appear on the student website

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