

Ch. 11 Gas Laws Chapter Review Answers

1. Particles of matter are always in motion. Gas particles are small and have a huge amount of space between them compared to their size. Because they are so far apart, they have no attraction between each other, and they have highly elastic collisions. The temperature of the gas is the average kinetic energy of the gas particles.
2. An ideal gas is a theoretical gas that has perfect particle qualities that fit the molecular-kinetic theory of matter.
3. Pressure:
 - a. The fast moving gas particles strike the sides of the container they are in.
 - b. The same force applied to a smaller area will calculate to a greater pressure.
4. Mercury and Pressure:
 - a. The pressure of the mercury in the tube is equal to the atmospheric pressure at sea-level, which is 101.325 kPa, or 1 atm, or 760 mmHg.
 - b. 10.3 m, or approximately 34 feet.
 - c. Mercury's density is 13.5 times greater than the density of water, therefore, it exerts 13.5 times more pressure than a column of water the same size.
5. Units of Pressure:
 - a. mmHg, in.Hg, torr, bar, psi, atm., Pa
 - b. 1 atm = 760 mmHg
 - c. A Pascal is the force of 1 N exerted over an area of 1 m². 1 Pa = 1N/1 m².
 - d. 1 atm = 101.325 kPa, or 101,325 Pa
6. Dalton's Law of Partial Pressures
 - a. The partial pressures are the pressures that each individual gas exerts in the mixture of gases.
 - b. They are independent of each other.
7. 543 mm
8. Torr units
 - a. 950. torr
 - b. 1.88 torr
 - c. 3.61×10^7 torr
 - d. 5.78×10^9 torr
9. More pressure units
 - a. 0.164 atm
 - b. 324,000 Pa or 3.24×10^5 Pa
 - c. 40.4 mmHg
10. 166.190 torr
11. 699.8 torr
12. Inverse Proportion
13. The molecules are closer together to each other and the container walls, and strike a given unit area of the container of the walls of the container more frequently.
14. Direct proportion
15. Direct proportion
16. Gas particles that are at a higher temperature will have a greater kinetic energy. Therefore at a constant volume for the container, the gas particles will collide with a given unit area of the walls of the container with more force and more frequently resulting in a greater pressure.
17. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$ Or $\frac{PV}{T} = k$
18. Gases!
 - a. 100.0 mL
 - b. 1.4 mmHg
19. Calculations:
 - a. 93.3 mL

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DP Hallahan



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