

Name: _____

KEY**Honors Chemistry Quick-Check: The Periodic Table**

- 101 Md Element named in honor of the "Father of the Periodic Table"
- number The chemical and physical properties are a function of their atomic (mass / number).
- period Term for horizontal rows
- family group Term for vertical columns
- p Element of 3rd period, $\frac{1}{2}$ filled p sublevel.
- Mg Which element(s) has no d electrons? Mg, Au, I, Cs
- 1, 0, Mg Number of elements in 3rd period with only 2 valence electrons
- Halogens Family group that ends with s^2p^5
- Noble Gas Family with filled valence shell
- He Element in family with filled valence shell that only contains s electrons.
- ~~Fe~~ Sb $[\text{Kr}]5s^24d105p^3$
- Ca 4th period alkaline Earth metal
- 1 Number of valence electrons that the alkali metals have
- Li lightest alkali metal
- Fr largest alkali metal
- Na largest element in 3rd period.
- $4f^2, 3p^2, 2d^1$ Which electron configuration(s) is/are not allowed: $7p^2, 4f^2, 3p^2, 2d^1$
- $3, 0, 1, -\frac{1}{2}$ 2 electrons in an orbital. If quantum set = $3, 0, 1, +\frac{1}{2}$ what is other quantum set of numbers?
- Te Ground state, 5th period element with valence orbital notation: $\uparrow\downarrow \quad \uparrow\downarrow \quad \uparrow \quad \uparrow$
- Sf Sublevel filling in the actinide series
- Bi element in 6th period with 5 valence electrons
- 10 maximum number of electrons allowed in the 3d sublevel.
- 12 How many s sublevel electrons does an atom of lead have? ($1s^2, 2s^2, 3s^2, 4s^2, 5s^2, 6s^2$)
- 2 6 4 Number of valence electrons for Sr, S, Pb
- 16 Number of orbitals for the 4th energy level $2n^2 = 2(4^2) = 32 e^- / 2 = 16 \text{ orbitals}$
- Pauli Exclusion P. No 2 electrons in the same atom can have the same 4 quantum numbers.

2. Circle the larger particle

- Mg or Ca
- P or P³⁻
- K or As
- Sn or S
- Al^{3+} or Al

- Ni or Ag
- I or P
- O^{2-} or S^{2-}
- Ca^{2+} or Cl⁻
- As or Br

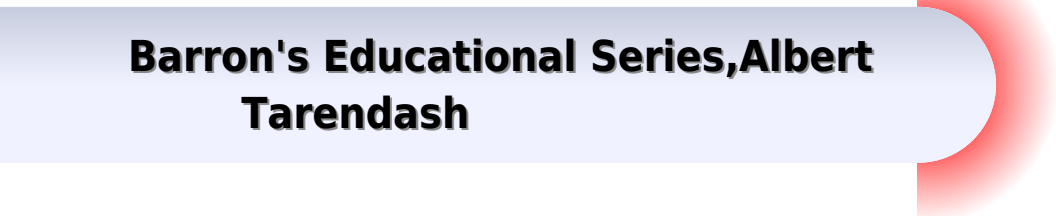
3. Circle particle with larger ionization energy

- Se or S
- K or Cs
- F or O

- O or P
- Al or Ca
- I or O

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