

Example:

25. Three grams of carbon and 15 g of oxygen are placed in a closed vessel and heated to 1000°C. What mass of carbon dioxide is formed?

$$25 \text{ g } C + \frac{1 \text{ mol } C}{12 \text{ g } C} = \frac{2}{3} \text{ mol } C \quad \frac{15 \text{ g } O_2}{32 \text{ g } O_2} = \frac{3}{8} \text{ mol } O_2 \quad \frac{3 \text{ mol } CO_2}{1 \text{ mol } O_2} = \boxed{7.5 \text{ g } CO_2}$$

26. When 10 g of Al reacts with an excess of Fe_2O_3 , 20 g of Al_2O_3 are produced. What is the percentage yield of this reaction?



$$10 \text{ g } Al = \frac{1 \text{ mol } Al}{27 \text{ g } Al} = \frac{2 \text{ mol } Al_2O_3}{2 \text{ mol } Al} = \frac{10 \text{ g } Al_2O_3}{1 \text{ mol } Al_2O_3} = 50.6 \text{ g } Al_2O_3 \text{ theoretical yield}$$

$$\% \text{ yield} = \frac{20 \text{ g } Al_2O_3}{50.6 \text{ g } Al_2O_3} \times 100 = \boxed{39.5\%}$$

27. Following the problem 26, what is the limiting reagent when 10 g of Al and 10 g of Fe_2O_3 react with 10 g of Al?

$$10 \text{ g } Al_2O_3 = \frac{1 \text{ mol } Al_2O_3}{102 \text{ g } Al_2O_3} = \frac{3 \text{ mol } Al}{2 \text{ mol } Al_2O_3} = \frac{2 \text{ g } Al}{1 \text{ mol } Al} = 3.0 \text{ g } Al$$

$$25 \text{ g } Fe_2O_3 = \frac{1 \text{ mol } Fe_2O_3}{160 \text{ g } Fe_2O_3} = \frac{3 \text{ mol } Al}{2 \text{ mol } Fe_2O_3} = \frac{2 \text{ g } Al}{1 \text{ mol } Al} = \boxed{2.75 \text{ g } Al} \text{ maximum product}$$

Limiting Reagent = Fe_2O_3

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Puja Mehta



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