

11. In the reaction: $4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$, if 20 grams of NH_3 and 50 grams of O_2 are mixed, which is the limiting reactant, and how many grams of NO can be produced? (Molar masses: $\text{NH}_3 = 17 \text{ g/mol}$, $\text{O}_2 = 32 \text{ g/mol}$, $\text{NO} = 30 \text{ g/mol}$)
12. For the reaction: $2\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{H}_2$, if 10 grams of Al and 50 grams of H_2SO_4 are reacted, which reactant is limiting, and how many grams of H_2 can be produced? (Molar masses: $\text{Al} = 27 \text{ g/mol}$, $\text{H}_2\text{SO}_4 = 98 \text{ g/mol}$, $\text{H}_2 = 2 \text{ g/mol}$)
13. In the reaction: $3\text{CuSO}_4 + 2\text{Na}_3\text{PO}_4 \rightarrow \text{Cu}_3(\text{PO}_4)_2 + 3\text{Na}_2\text{SO}_4$, if 30 grams of CuSO_4 and 20 grams of Na_3PO_4 are mixed, which is the limiting reactant, and how many grams of $\text{Cu}_3(\text{PO}_4)_2$ can be produced? (Molar masses: $\text{CuSO}_4 = 160 \text{ g/mol}$, $\text{Na}_3\text{PO}_4 = 164 \text{ g/mol}$, $\text{Cu}_3(\text{PO}_4)_2 = 382 \text{ g/mol}$)
14. For the reaction: $2\text{KMnO}_4 + 5\text{H}_2\text{C}_2\text{O}_4 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{MnSO}_4 + 10\text{CO}_2 + 8\text{H}_2\text{O} + \text{K}_2\text{SO}_4$, if 20 grams of KMnO_4 , 30 grams of $\text{H}_2\text{C}_2\text{O}_4$, and 25 grams of H_2SO_4 are reacted, which reactant is limiting, and how many grams of CO_2 can be produced? (Molar masses: $\text{KMnO}_4 = 158 \text{ g/mol}$, $\text{H}_2\text{C}_2\text{O}_4 = 90 \text{ g/mol}$, $\text{H}_2\text{SO}_4 = 98 \text{ g/mol}$, $\text{CO}_2 = 44 \text{ g/mol}$)

Holt Chemistry Limiting Reactants And Percentage Yield

Neil D. Jespersen, Alison Hyslop



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