

HOLT MODERN CHEMISTRY – CHAPTER 1 REVIEW

The following pages contain the bulk (but not all) of the information for the chapter 1 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test.
****** Test items will be recall, examples, and/or application of this content. ******

OUTCOMES

- 1.1: Distinguish between types of research – basic, applied, and technological development. (F & PK)
- 1.2: Distinguish between the physical and chemical properties of matter. (F & PK)
- 1.2: Distinguish among homogeneous and heterogeneous mixtures, elements, and compounds. (F & PK)
- 1.3: List properties of metals and non-metals. (T & R)
- 2.3: Determine significant figures. (F & PK)

1.1: CHEMISTRY IS A PHYSICAL SCIENCE

- **Vocabulary**
 - **chemistry** – the scientific study of the composition, structure, and properties of matter and the changes that matter undergoes
 - **chemical** – any substance with a defined composition
- **Chapter Highlights**
 - Chemistry is the study of the composition, structure, and properties of matter and the changes that matter undergoes.
 - A chemical is any substance that has a definite composition or is used or produced in a chemical process.
 - Basic research is carried out for the sake of increasing knowledge. Applied research is carried out to solve practical problems. Technological development involves the use of existing knowledge to make life easier or more convenient.

1.2: MATTER AND ITS PROPERTIES

- **Vocabulary**
 - **matter** – anything that has mass and takes up space
 - **mass** -- a measure of the amount of matter in an object
 - **element** – a substance that cannot be separated or broken down into simpler substances by chemical means; all atoms of an element have the same atomic number
 - **atom** – the smallest unit of an element that maintains the chemical properties of that element
 - **compound** -- a substance made up of atoms of two or more different elements joined by chemical bonds
 - **extensive property** -- a property that depends on the extent or size of a system
 - **intensive property** -- a property that does not depend on the amount of matter present, such as pressure, temperature, or density
 - **physical property** -- a characteristic of a substance that does not involve a chemical change, such as density, color, or hardness
 - **physical change** -- a change of matter from one form to another without a change in chemical properties
 - **change of state** -- the change of a substance from one physical state to another
 - **solid** -- the state of matter in which the volume and shape of a substance are fixed
 - **liquid** -- the state of matter that has a definite volume but not a definite shape
 - **gas** -- a form of matter that does not have a definite volume or shape
 - **plasma** -- in physical science, a state of matter that starts as a gas and then becomes ionized; it consists of freemoving ions and electrons and it takes on an electric charge
 - **chemical property** -- a property of matter that describes a substance's ability to participate in chemical reactions
 - **chemical change** -- a change that occurs when one or more substances change into entirely new substances with different properties
 - **chemical reaction** -- the process by which one or more substances change to produce one or more different substances
 - **reactant** -- a substance or molecule that participates in a chemical reaction
 - **product** -- a substance that forms in a chemical reaction

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web stoichiometric principles can be used to determine the chemical equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction the iron ions dissolve in the copper ii sulfate solution to produce copper ii ions

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web lab 4 stoichiometry the reaction of iron with copper ii sulfate revised 8 19 2009 1 introduction in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces

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web under the finalization of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry lab report the reaction of iron with copper ii sulfate period 5 in this experiment i observed the reaction of iron with copper ii sulfate i obtained the results to this experiment through adding heated CuSO_4 solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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web stoichiometry the reaction of iron with copper ii sulfate purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction

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