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John Haydon



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determine based on electronegativity and symmetry if the following molecules are polar or nonpolar HCN $\text{H}_2\text{C}_2\text{F}_4$ COCl_2 COCl_2 NF_3 NF_3 Q_3

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web the relative electronegativities of the bonded atoms is $\text{H} < \text{C} < \text{Cl}$ and so the bond moments all point toward the Cl end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as BF_3 trigonal planar CH_4 tetrahedral PF_5 trigonal bipyramidal and SF_6

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