

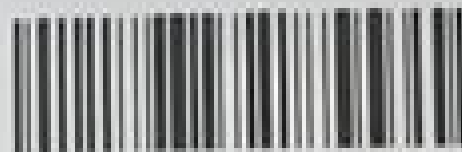


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$$\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$$

$$2\text{C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

$$3\text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} \rightarrow \text{KBr} + \text{CO}_2 + \text{NaBr} + \text{NaNO}_3$$
 using the equation from problem 2 above answer the following

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$$2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$3\text{HBr} + 1\text{KHCO}_3 + 1\text{H}_2\text{O} \rightarrow 1\text{KBr} + 1\text{CO}_2 + 2\text{NaBr} + 3\text{NaNO}_3$$
 using the following equation

$$2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$$

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$$\text{Ce} + \text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu} + \text{Zn}$$
 how many moles of Ce will be produced from 23.0 g of Ce assuming Ce is available in excess

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$$\text{Br}_2 + \text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$$

$$2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$\text{Br}_2 + 3\text{HBr} + 1\text{KHCO}_3 + 1\text{H}_2\text{O} \rightarrow 1\text{KBr} + 1\text{CO}_2 + 2\text{NaBr} + 3\text{NaNO}_3$$
 using the following equation

$$\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$$

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