

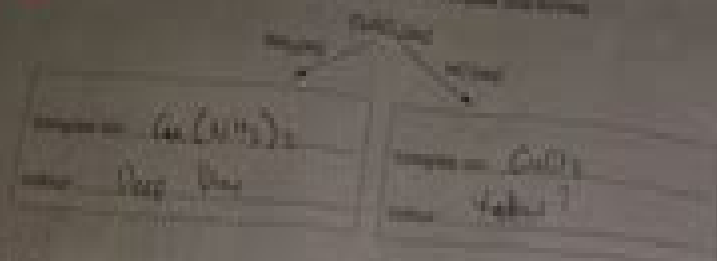
Question 1

Answer all the questions

The reaction is about the synthesis of copper complexes and ligands.

(a) The reaction shows the synthesis of a complex compound.

In the boxes, write the formula and name of the complex and ligand.



(b) Cu^{2+} ions form a complex ion B with two ethylenediamine ligands and two water molecules. The ethylenediamine ion is a bidentate ligand.

The chemical formula of the ethylenediamine ion is shown in Fig. 1.1 below.



Fig. 1.1

(c) What is meant by the term bidentate ligand?

A ligand that has two donor atoms, each of which can form a coordinate bond with the central metal ion.

(d) The complex ion A is shown in Fig. 1.2 below.

The structure of one of the ethylenediamine ligands is shown below. The charge has been given.



Complex A

Complete the 3D diagram of the other two geometries of A. You do not need to include any charges.

Indicate with ticks whether the stereoisomers are cis, trans, optical or a combination of these types.

Stereoisomer														
Type	<table><tr><td>cis</td><td></td></tr><tr><td>trans</td><td>✓</td></tr><tr><td>optical</td><td></td></tr></table>	cis		trans	✓	optical		<table><tr><td>cis</td><td>✓</td></tr><tr><td>trans</td><td></td></tr><tr><td>optical</td><td>✓</td></tr></table>	cis	✓	trans		optical	✓
cis														
trans	✓													
optical														
cis	✓													
trans														
optical	✓													

(e) What is the empirical formula, including the charge, of the complex ion A?

$[Cu(NH_3)_2(en)_2]^{2+}$

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Lauren Gardner



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