

F325 Acids, Bases and Buffers test ms

1. (a) completely dissociates/ionised (1) 2
proton donor (1)
- (b) NO_3^- (1) 1
- (c) (i) $\text{pH} = -\log[\text{H}^+] / -\log(0.015)$ (1) = 1.82 / 1.8 (1) (Not 2) 2
(ii) $[\text{H}^+] = 0.0075 \text{ mol dm}^{-3}$
 $\text{pH} = -\log(0.0075) = 2.12 / 2.1$ (1) 1
- [6]**
2. (a) (i) H^+ /proton donor (1) 1
(ii) partially dissociates/ionises (1) 1
- (b) $\text{C}_6\text{H}_5\text{OH}(\text{aq}) + \text{OH}^-(\text{aq}) \rightleftharpoons \text{C}_6\text{H}_5\text{O}^-(\text{aq}) + \text{H}_2\text{O}(\text{l})$
acid 1 base 2 (1) base 1 acid 2 (1)
1 mark for each acid-base pair
- (c) (i) $K_a = \frac{[\text{C}_6\text{H}_5\text{O}^-][\text{H}^+]}{[\text{C}_6\text{H}_5\text{OH}]}$ (1) 1
(ii) concentration = $38/94$ (1) = 0.40 mol dm^{-3} (1)
(first mark for M_r of phenol – incorrect answer here will give ecf for remainder of question)
- $$1.3 \times 10^{-10} \approx \frac{[\text{H}^+(\text{aq})]^2}{0.40} \quad (1)$$
- (‘=’ sign is acceptable)
- $[\text{H}^+] = \sqrt{((1.3 \times 10^{-10}) \times (0.40))} = 7.2 \times 10^{-6} \text{ mol dm}^{-3}$ (1)
 $\text{pH} = -\log[\text{H}^+] = -\log 7.2 \times 10^{-6} = 5.14$ (1) 5
- 3 marks: $[\text{H}^+]$ (1), pH expression (1), calc of pH from $[\text{H}^+]$ (1)
- Common errors:**
Without square root, answer = 10.28 (1)(1)(0)
Use of 38 as molar concentration does not score 1st 2 marks.
This gives an answer of 4.15 for 3 marks (1)(1)(1)
- (d) $[\text{H}^+(\text{aq})] = 1.99 \times 10^{-6} \text{ mol dm}^{-3}$ (1)
 $[\text{C}_6\text{H}_5\text{O}^-(\text{aq})] = K_a [\text{C}_6\text{H}_5\text{OH}(\text{aq})] / [\text{H}^+(\text{aq})]$ (1)
 $[\text{C}_6\text{H}_5\text{O}^-(\text{aq})] = 0.013 \text{ mol dm}^{-3}$ (1)
- Calculation should use half the original concentration of phenol to find the concentration of sodium phenoxide in the buffer. This should then be doubled back up again.*
Do not penalise an approach that uses the original concentration of phenol in the expression above.

[13]

F325 Synoptic Questios

Rachel S Tattersall



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Table of Contents F325 Synoptic Questios

1. Understanding the eBook F325 Synoptic Questios
 - The Rise of Digital Reading F325 Synoptic Questios
 - Advantages of eBooks Over Traditional Books
2. Identifying F325 Synoptic Questios
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an F325 Synoptic Questios
 - User-Friendly Interface
4. Exploring eBook Recommendations from F325 Synoptic Questios
 - Personalized Recommendations
 - F325 Synoptic Questios User Reviews and Ratings
 - F325 Synoptic Questios and Bestseller Lists
5. Accessing F325 Synoptic Questios Free and Paid eBooks
 - F325 Synoptic Questios Public Domain eBooks
 - F325 Synoptic Questios eBook Subscription Services
 - F325 Synoptic Questios Budget-Friendly Options

6. Navigating F325 Synoptic Questios eBook Formats
 - ePub, PDF, MOBI, and More
 - F325 Synoptic Questios Compatibility with Devices
 - F325 Synoptic Questios Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of F325 Synoptic Questios
 - Highlighting and Note-Taking F325 Synoptic Questios
 - Interactive Elements F325 Synoptic Questios
8. Staying Engaged with F325 Synoptic Questios
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers F325 Synoptic Questios
9. Balancing eBooks and Physical Books F325 Synoptic Questios
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection F325 Synoptic Questios
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine F325 Synoptic Questios
 - Setting Reading Goals F325 Synoptic Questios
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of F325 Synoptic Questios
 - Fact-Checking eBook Content of F325 Synoptic Questios
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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