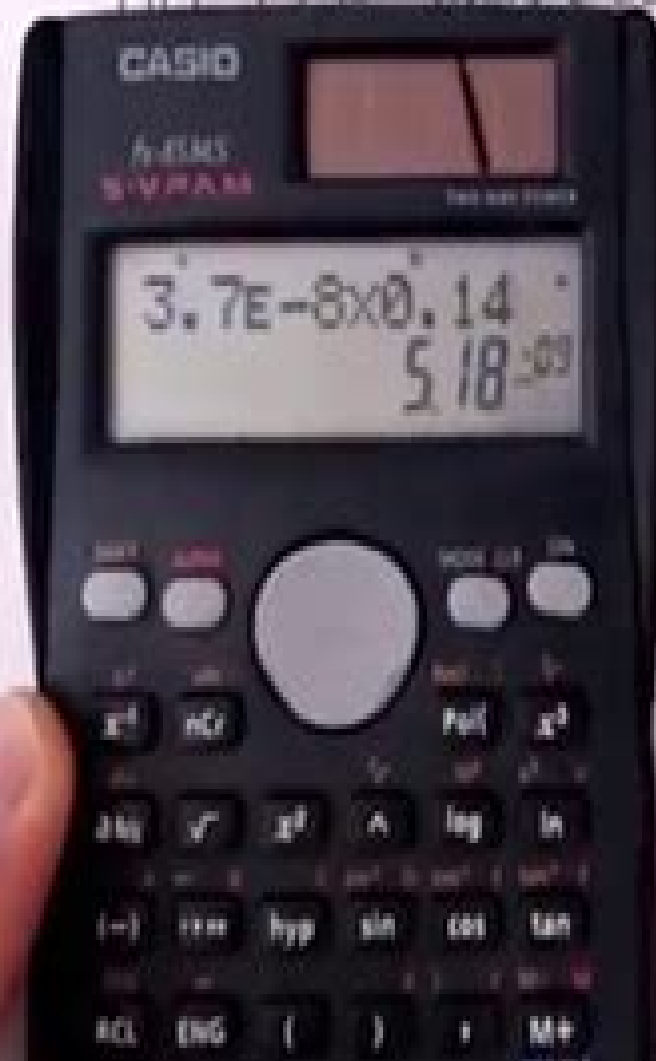


- (a) 0.14 mol dm^{-3} solutions of hydrochloric acid, HCl , and chloroic acid, HClO ($K_a = 1.7 \times 10^{-8}$), have different pH values.

Explain why the pH values are different and calculate the pH of 0.14 mol dm^{-3} solutions of HCl and HClO to two decimal places.

Show any working in calculations.

HCl is a strong acid and HClO is



$$\rightarrow -\log 0.14 = 0.85.$$

$$7 \times 10^{-8}$$

$$[\text{H}^+]^2 = 3.7 \times 10^{-8} \times 0.14$$

$$= 5.18 \times 10^{-9}$$

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Huangqi Zhang



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Table of Contents F325 June 2013 Past Paper Pack

1. Understanding the eBook F325 June 2013 Past Paper Pack
 - The Rise of Digital Reading F325 June 2013 Past Paper Pack
 - Advantages of eBooks Over Traditional Books
2. Identifying F325 June 2013 Past Paper Pack
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an F325 June 2013 Past Paper Pack
 - User-Friendly Interface
4. Exploring eBook Recommendations from F325 June 2013 Past Paper Pack
 - Personalized Recommendations
 - F325 June 2013 Past Paper Pack User Reviews and Ratings
 - F325 June 2013 Past Paper Pack and Bestseller Lists
5. Accessing F325 June 2013 Past Paper Pack Free and Paid eBooks
 - F325 June 2013 Past Paper Pack Public Domain eBooks
 - F325 June 2013 Past Paper Pack eBook Subscription Services
 - F325 June 2013 Past Paper Pack Budget-Friendly Options
6. Navigating F325 June 2013 Past Paper Pack eBook Formats
 - ePub, PDF, MOBI, and More
 - F325 June 2013 Past Paper Pack Compatibility with Devices
 - F325 June 2013 Past Paper Pack Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of F325 June 2013 Past Paper Pack
 - Highlighting and Note-Taking F325 June 2013 Past Paper Pack
 - Interactive Elements F325 June 2013 Past Paper Pack
8. Staying Engaged with F325 June 2013 Past Paper Pack

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers F325 June 2013 Past Paper Pack
9. Balancing eBooks and Physical Books F325 June 2013 Past Paper Pack
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection F325 June 2013 Past Paper Pack
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine F325 June 2013 Past Paper Pack
- Setting Reading Goals F325 June 2013 Past Paper Pack
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of F325 June 2013 Past Paper Pack
- Fact-Checking eBook Content of F325 June 2013 Past Paper Pack
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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