

Question	Marking Guidance	Mark	Comments
1(a)	<p><u>Enthalpy change</u> for the formation of <u>1 mol</u> of <u>gaseous atoms</u></p> <p>From the <u>element</u> (in its standard state)</p> <p>Enthalpy change to separate <u>1 mol</u> of an <u>ionic</u> lattice/solid/compound</p> <p>Into (its component) <u>gaseous ions</u></p>	<p>1</p> <p>1</p> <p>1</p> <p>1</p>	<p>allow <u>heat energy change</u> for <u>enthalpy change</u></p> <p>ignore reference to conditions</p> <p>enthalpy change not required but penalise energy</p> <p>mark all points independently</p>
1(b)	<p><math>\Delta H_f^\circ = \Delta H_f^\circ + \Delta H_{at}^\circ + \text{I.E.} + 1/2E(\text{Cl}-\text{Cl}) + \text{EA}</math></p> <p><math>= +411 + 109 + 494 + 121 - 364</math></p> <p><math>= +771 \text{ (kJ mol}^{-1}\text{)}</math></p>	<p>1</p> <p>1</p> <p>1</p>	<p>Or correct Born-Haber cycle drawn out</p> <p>–771 scores 2/3</p> <p>+892 scores 1/3</p> <p>–51 scores 1/3</p> <p>–892 scores zero</p> <p>+51 scores zero ignore units</p>
1(c)(i)	<p>Ions are perfect spheres (or point charges)</p> <p><u>Only</u> electrostatic attraction/no covalent interaction</p>	<p>1</p> <p>1</p>	<p>mention of molecules/intermolecular forces/covalent bonds CE = 0</p> <p>allow ionic bonding <u>only</u></p> <p>If mention of atoms CE = 0 for M2</p>
1(c)(ii)	ionic	1	Allow no covalent character/bonding

# January 2019aqa His11 Mark Scheme

**Roman Wölfel**



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