

Half-Modern Chemistry Review

CHAPTER 11: CHEMICAL BONDING

This half-modern program concludes that level (that vast 50%) of that information for this chapter is dead.

Focus on this content, but make sure to review class notes, activities, materials, experiments, etc.

If you study this document and DISCOVER what, you should at least be able to PLUG that fact.

Now that there will be recall, examples, and/or application of this content, now

OUTCOMES

- Communicate with precision the understanding/properties learned (p2 c2)
- Communicate understanding learned to teachers and/or peers (p2 c2)
- EL-2 Give the Lewis dot notation, illustrate the placement bond on the drawing of molecules (p2 d, p2)
- EL-2 List important physical properties of some compounds; name and illustrate molecules in ions (p2 d, p2)

EL-1: INTRODUCTION TO CHEMICAL BONDING

• Introductory

- a) chemical bond → the attraction force that binds atoms or ions together
- a) ionic bonding → a force that attracts particles that have been attracted, which demonstrates a mutual attraction, and can be
- a) covalent bonding → a bond formed when atoms share electrons on or around points of total bond.
- a) intermolecular bonding → a bond formed by attractive forces between atoms and molecules attracted to bonds between atoms.
- a) polar → substance, a molecule or particle that particles and charged ions charge, and comparison
- a) polar covalent bond → a chemical bond in which a pair of electrons shared by two atoms is held more closely to one atom.

• Chapter Highlights

- a) Many atoms are chemically bonded to other atoms. The most major types of chemical bonding are ionic, covalent, and metallic.
- a) In general, atoms of metals, bond covalently with atoms of nonmetals, atoms of metals bond covalently with each other, and atoms of nonmetals bond covalently with each other.

EL-2: COVALENT BONDING AND MOLECULAR COMPOUNDS

• Introductory

- a) molecule → a group of atoms that are held together by chemical bonds; a molecule is the smallest unit of matter that can exist by itself and retain all of its chemical properties
- a) molecular compound → a chemical compound whose simplest unit, and represents
- a) structural formula → a representation of chemical formula, and illustrates the arrangement of molecules
- a) molecular formula → a chemical formula that shows the number and kinds of atoms in a molecule, but not the arrangement of the atoms.

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Jicheng Xie

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