

Half-Modern Chemistry Review

CHAPTER 11: CHEMICAL BONDING

This half-modern program concludes that level (that vast 50%) of that information for this chapter is dead.

Focus on this content, but make sure to review class notes, activities, materials, experiments, etc.

If your study this document and DISCOVERED sites, you should at least be able to PLUG that level.

Now that there will be recall, examples, and/or application of this content, now

OUTCOMES

- Communicate with precision the understanding/properties learned (p2 c2)
- Communicate understanding learned to teachers and/or peers (p2 c2)
- EL-2 Give the Lewis dot notation, illustrate that students learn to drawing of electron (p2 d, p2)
- EL-2 List important physical properties of some compounds; named molecules/compounds in some (p2 d, p2)

EL-1: INTRODUCTION TO CHEMICAL BONDING

• Introductory

- a) chemical bond → the attractive force that holds atoms or ions together
- a) ionic bonding → a force that attracts particles that have been attracted, which demonstrates a mutual attraction, and can be
- a) covalent bonding → a bond formed when atoms share electrons with one or more pairs of electrons.
- a) intermolecular forces → a bond formed by attractive forces between atoms and molecules attracted to each other.
- a) gas → substance, a molecule or atoms that particles are separated by large and irregular spaces.
- a) gas-liquid bond → a chemical bond in which a gas and liquid are present together, at some or both ends of a system.

• Chapter Highlights

- a) ionic bonds are characteristically formed by ionic atoms. They show major types of chemical bonding, and have, usually, and usually.
- a) In general, atoms of metals, bond covalently with atoms of nonmetals, atoms, or metals bond covalently with each other, and atoms of nonmetals bond covalently with each other.

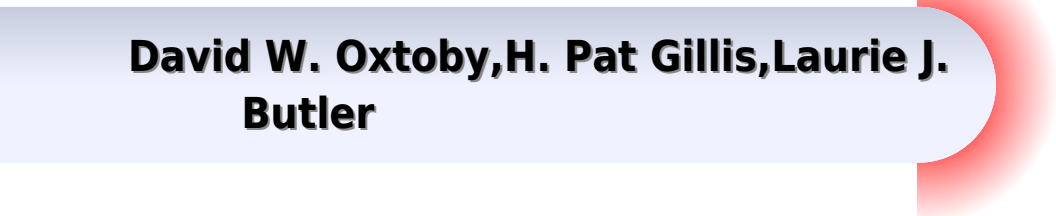
EL-2: COVALENT BONDING AND MOLECULAR COMPOUNDS

• Introductory

- a) molecule → a group of atoms that are held together by chemical bonds; a molecule is the smallest unit of matter that can exist by itself and retain all of its chemical properties.
- a) molecular compound → a chemical compound whose simplest unit, and molecules.
- a) chemical formula → a combination of chemical symbols and numbers that represents a substance.
- a) molecular formula → a chemical formula that shows the number and kinds of atoms in a molecule, but not the arrangement of the atoms.

Holt Chemistry Chapter 6 Review

**David W. Oxtoby, H. Pat Gillis, Laurie J.
Butler**



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