

→ CHEMICAL EQUILIBRIUM

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Q.1: Justify that chemical equilibrium is dynamic in nature?

Ans: In reversible chemical reactions the molecules of reactions collide and convert into products. At the same time the molecules of the products are converted into reactants. When two opposing forces maintain the equal, rates then equilibrium is there and that is dynamic equilibrium in nature.

Q.2: When does four types of chemical equilibrium constants for a reaction become equal?

Ans: When we have such a reversible reaction, in which the total number of moles of reactants and products are equal then $K_p = K_c = K_n$. In other words $\Delta n = 0$, where Δn is the difference of number of moles of reactants and products.

Q.3: The change of temperature disturbs the equilibrium position and the equilibrium constant of reaction? Justify?

Ans: All the reversible reactions are disturbed by changing their equilibrium position and equilibrium constant by disturbing the temperature. Actually change of temperature changes the energy contents of reactants and products.

Q.4: Why is solubility of glucose in water is increased by increasing the temperature?

Ans: The solubility of glucose in water is an endothermic process. Increase of temperature pushed the system to that side where heat is absorbed, and that is the side where solubility increases.

Q.5: What is the nature of solution having $\text{pH} = 7.5$?

Ans: The solution with $\text{pH} = 7$ is neutral and having values more than 7 have greater concentration of OH^- and less concentration of H^+ , so the solution is basic.

Q.6: What is buffer capacity?

Ans: Buffer capacity is a ability of a buffer to resist the change of pH , when a few drops of an acid and base is added from outside. The best buffer is obtained. (i) when pH of the buffer is equal to pK_a of the acid and the concentration of the salt and acid are equal. (ii) The OH^- of the buffer is equal to pK_b or the base.

Holt Chemistry Concept Review Answers Chemical Equilibrium

Mark Steyvers



Holt Chemistry Concept Review Answers Chemical Equilibrium:

This book delves into Holt Chemistry Concept Review Answers Chemical Equilibrium. Holt Chemistry Concept Review Answers Chemical Equilibrium is a crucial topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Holt Chemistry Concept Review Answers Chemical Equilibrium, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
 - Chapter 1: Introduction to Holt Chemistry Concept Review Answers Chemical Equilibrium
 - Chapter 2: Essential Elements of Holt Chemistry Concept Review Answers Chemical Equilibrium
 - Chapter 3: Holt Chemistry Concept Review Answers Chemical Equilibrium in Everyday Life
 - Chapter 4: Holt Chemistry Concept Review Answers Chemical Equilibrium in Specific Contexts
 - Chapter 5: Conclusion
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 3. In chapter 2, the author will delve into the foundational concepts of Holt Chemistry Concept Review Answers Chemical Equilibrium. The second chapter will elucidate the essential principles that must be understood to grasp Holt Chemistry Concept Review Answers Chemical Equilibrium in its entirety.
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 6. In chapter 5, the author will draw a conclusion about Holt Chemistry Concept Review Answers Chemical Equilibrium. This chapter will summarize the key points that have been discussed throughout the book.
- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Holt Chemistry Concept Review Answers Chemical Equilibrium.

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