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Acid-Base Theories

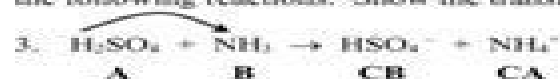
1. Summarize the three main acid-base theories in the table below.

	ACID	BASE
Arrhenius	H^+ or H_3O^+ producer in aqueous solution	OH^- producer in aqueous solution
Bronsted-Lowry	proton (H^+) donor	proton (H^+) acceptor
Lewis	electron-pair acceptor	electron-pair donor

Use the following chemical equations to answer question 2.



2. According to Arrhenius, HCl would be classified as a(n)
- acid
- and
- NH_3
- would be classified as a(n)
- base
- . Explain why.

HCl produces hydronium ions (H_3O^+) in water. NH_3 produces hydroxide ions (OH^-) in water.For problems 3 – 5, label the acid (A), base (B), conjugate acid (CA), and conjugate base (CB) in each of the following reactions. Show the transfer of the proton (H^+) by drawing an arrow.

For problems 6 – 9, give the conjugate base for each Bronsted-Lowry acid.



For problems 10 – 13, give the conjugate acid for each Bronsted-Lowry base.



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1. Understanding the eBook Holt Chemistry Acid And Bases Review Answers
 - The Rise of Digital Reading Holt Chemistry Acid And Bases Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Chemistry Acid And Bases Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Chemistry Acid And Bases Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Chemistry Acid And Bases Review Answers
 - Personalized Recommendations
 - Holt Chemistry Acid And Bases Review Answers User Reviews and Ratings
 - Holt Chemistry Acid And Bases Review Answers and Bestseller Lists

5. Accessing Holt Chemistry Acid And Bases Review Answers Free and Paid eBooks
 - Holt Chemistry Acid And Bases Review Answers Public Domain eBooks
 - Holt Chemistry Acid And Bases Review Answers eBook Subscription Services
 - Holt Chemistry Acid And Bases Review Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Holt Chemistry Acid And Bases Review Answers Compatibility with Devices
 - Holt Chemistry Acid And Bases Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Chemistry Acid And Bases Review Answers
 - Highlighting and Note-Taking Holt Chemistry Acid And Bases Review Answers
 - Interactive Elements Holt Chemistry Acid And Bases Review Answers
8. Staying Engaged with Holt Chemistry Acid And Bases Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Chemistry Acid And Bases Review Answers
9. Balancing eBooks and Physical Books Holt Chemistry Acid And Bases Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Chemistry Acid And Bases Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Holt Chemistry Acid And Bases Review Answers
 - Setting Reading Goals Holt Chemistry Acid And Bases Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Holt Chemistry Acid And Bases Review Answers
 - Fact-Checking eBook Content of Holt Chemistry Acid And Bases Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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