#### Chapter 19 Electrochemistry

#### Practice questions. Answers

What element is being reduced in the following redox reaction?

$$H_2O_2(l) + ClO_2(aq) \rightarrow ClO_2^-(aq) + O_2(g)$$
 Cl

2. What element is being oxidized in the following redox reaction?

$$H_2O_2(l) + ClO_2(aq) \rightarrow ClO_2^-(aq) + O_2(g)$$

3. What element is being oxidized in the following redox reaction?

$$C3H8O2(aq) + KMnO4(aq) \rightarrow C3H2O4K2(aq) + MnO2(aq) C$$

4. What element is being reduced in the following redox reaction?

$$C_3H_8O_2(aq) + KMnO_4(aq) \rightarrow C_3H_2O_4K_2(aq) + MnO_2(aq) Mn$$

Balance the following redox reaction if it occurs in acidic solution. What are the coefficients in front
of H2C2O4 and H2O in the balanced reaction?

$$MnO_4^-(aq) + H_2C_2O_4(aq) \rightarrow Mn^2+(aq) + CO_2(g)$$

#### $H_2C_2O_4 = 5$ , $H_2O = 8$

6. Balance the following redox reaction if it occurs in acidic solution. What are the coefficients in front of H\* and Fe3+ in the balanced reaction?

$$Fe^{2+}(aq) + MnO_4^-(aq) \rightarrow Fe^{3+}(aq) + Mn^{2+}(aq)$$

$$H^* = 8$$
,  $Fe^{3+} = 5$ 

7. Determine the cell notation for the redox reaction given below.

$$3 \text{ Cl}_2(g) + 2 \text{ Fe(s)} \rightarrow 6 \text{ Cl}^-(aq) + 2 \text{ Fe}^{3+}(aq)$$

## Fe(s) | Fe3+(aq) | Cl2(g) | Cl-(aq) | Pt

 Use the standard half-cell potentials listed below to calculate the standard cell potential for the following reaction occurring in an electrochemical cell at 25°C. (The equation is balanced.)

$$3 \text{ Cl}_2(g) + 2 \text{ Fe}(s) \rightarrow 6 \text{ Cl}^-(aq) + 2 \text{ Fe}^{3+}(aq)$$

# **Fall 2012 Electrochemistry Test Answers**

**Daniel F McAuley** 

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