

## GROUP 2 ELEMENTS AND THEIR COMPOUNDS

- Show their outermost electron placed on s-subshell
- Usually need to lose 2 electrons
  - $Ca \rightarrow Ca^{2+} + 2e^{-}$

### Trends down the group

- $I^{\circ}$  ionisation energy decreases
  - Distance between outermost electron and nucleus increases
  - More shielding
  - Counteracts effect of increasing nuclear charge
- Atomic radius increases
- Melting point generally decreases
  - Much lower than d-transition metals

### Physical properties

- Good conductors of heat and electricity
  - Due to delocalised electrons
- Form white compounds
  - Transition metals form coloured compounds
- Oxidation number of +2
- All compounds are toxic to animal life
  - Be compounds show some covalent character because small highly charged cation polarises water shell of anion
- Oxides / hydroxides are basic
  - React with acids (proton acceptors)
  - Example:
    - $BeO$
    - $Be(OH)_2$  which is amphoteric
- Density =  $2-3.2 \text{ g/cm}^3$ 
  - Average density of transition metals is about  $7.5 \text{ g/cm}^3$
- Compared to group 2 metals:
  - Harder + denser
  - Higher melting point
    - Metallic bond is stronger (higher charge on cation & more electrons delocalised)

### Chemical properties

- Strong reducing agents
- Oxidation is losing electrons
- Reduction is gaining electrons
  - Reducing agent donates electrons (reducing other substances)
  - Oxidising agent accepts electrons (oxidising itself)
- E.g.
  - $2Ca + O_2 \rightarrow 2CaO$ 
    - $2Ca \rightarrow 2Ca^{2+} + 4e^{-}$
    - $O_2 + 4e^{-} \rightarrow 2O^{2-}$

# F321 May 2013 Ocr Chemistry Unofficial Paper

**D Keegan**



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