



Ford Ranger Manual Transmission Conversion

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produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have 0.216232 grams/mole 21.3 grams / 0.66 moles = 32.27 g

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web a write a balanced equation for the reaction $Al + O_2 \rightarrow Al_2O_3$ 4 Al + 3 $O_2 \rightarrow 2 Al_2O_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al x 1 mol Al / x 2 mol Al_2O_3 x 101.964 g Al_2O_3 / 302.3 g Al_2O_3 = 26.982 g Al 4 mol Al / 1 mol Al_2O_3 c determine the percent yield yield actual x 100 26.00 g / x 100 = 86.01% theoretical 302.3 g

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