

Name: _____ Date: _____ Period: _____

Midterm Exam Study Guide

d. Isotopes

- Isotopes differ in their _____ and number of _____.

e. Average atomic mass Formula:

f. Ions

- _____ have lost electrons and have a _____ charge
- _____ have gained electrons and have a _____ charge

g. Group Numbers be able to locate them

- Alkali Metals
- Alkaline earth Metals
- Transition Metals
- Inner Transition Metals
- Halogens
- Noble Gases

h. Periodic Trends Give your notes for reference (be able to for each of these)

- Atomic Radius/Size
- Ionic Radius/Size
- Ionization Energy
- Electronegativity

i. Valence Electrons be able to identify # of valence electrons and charge for each group

- Lewis Dot structures

F. Nuclear Decay and Half Life

- ${}^4_2\text{He}$ or ${}^4_2\text{He}$ is called an _____ particle which has the weakest radiation.
- ${}^0_{-1}\text{e}$ or ${}^0_{-1}\beta$ is called an _____ particle which has moderate radiation and is an electron.
- ${}^4_{-2}\text{He}$ is called an _____ particle which has the strongest radiation and highest energy.
- Half-life
 - Definition: _____
 - Know how to read a graph and calculate half-life

d. Electron Configurations

- Be able to draw them/identify them from another student
- Hund's Rule
- Pauli Exclusion Principle
- Aufbau Principle

e. Ionic Compounds

- Ionic compounds are made of _____ + _____
- Their overall charge is _____
- Know how to use Lewis Dot Structures to draw bonds between cation and anion

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