

15 Hydrocarbons

Alkanes

Chemical Reactivity

Alkanes do not react with many reagents.

This is because the C-C bond and the C-H bond are relatively strong and so cannot form stronger bonds with other elements that easily and there are no polar side groups.

Their main reactions are combustion reactions in air although they can react with halogens in the presence of UV light.

Combustion

Complete Combustion

In excess oxygen alkanes will burn with complete combustion.

The products of complete combustion are CO_2 and H_2O .



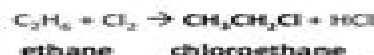
Alkanes readily burn in the presence of oxygen. This combustion of alkanes is highly exothermic, explaining their use as **fuels**.

Reaction of alkanes with bromine / chlorine in UV light

Synthesis of chloroalkanes

In the presence of **UV light** alkanes react with chlorine to form a mixture of products with the halogens substituting hydrogen atoms.

Overall Reaction



This is the overall reaction, but a more complex mixture of products is actually formed.

*To understand this reaction fully we must look in detail at how it proceeds step by step. This is called its **mechanism**.*

The **MECHANISM** for this reaction is called a **FREE RADICAL SUBSTITUTION**.

It proceeds via a series of steps:
STEP ONE: Initiation
STEP TWO: Propagation
STEP THREE: Termination

STEP ONE Initiation

Essential condition: UV light



The UV light supplies the energy to break the Cl-Cl bond. It is broken in preference to the others as it is the weakest.

The bond has broken in a process called **homolytic fission**.

*each atom gets **one** electron from the covalent bond*

When a bond breaks by homolytic fission it forms **Free Radicals**.

Free Radicals do not have a charge and are represented by a



DEFINITION

A **Free Radical** is a reactive species which possess an **unpaired electron**.

STEP TWO Propagation



The chlorine free radicals are very reactive and remove an H from the methane leaving a methyl free radical

The methyl free radical reacts with a Cl_2 molecule to produce the main product and another Cl free radical

All propagation steps have a **free radical** in the **reactants** and in the **products**.

As the Cl free radical is regenerated, it can react with several more alkane molecules in a **CHAIN REACTION**.

Hydrocarbons Chemistry Study Guide Answers

Barbara Ucko



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Table of Contents Hydrocarbons Chemistry Study Guide Answers

1. Understanding the eBook Hydrocarbons Chemistry Study Guide Answers
 - The Rise of Digital Reading Hydrocarbons Chemistry Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Hydrocarbons Chemistry Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hydrocarbons Chemistry Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hydrocarbons Chemistry Study Guide Answers
 - Personalized Recommendations
 - Hydrocarbons Chemistry Study Guide Answers User Reviews and Ratings
 - Hydrocarbons Chemistry Study Guide Answers and Bestseller Lists

5. Accessing Hydrocarbons Chemistry Study Guide Answers Free and Paid eBooks
 - Hydrocarbons Chemistry Study Guide Answers Public Domain eBooks
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 - Hydrocarbons Chemistry Study Guide Answers Budget-Friendly Options
6. Navigating Hydrocarbons Chemistry Study Guide Answers eBook Formats
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 - Hydrocarbons Chemistry Study Guide Answers Compatibility with Devices
 - Hydrocarbons Chemistry Study Guide Answers Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Hydrocarbons Chemistry Study Guide Answers
 - Highlighting and Note-Taking Hydrocarbons Chemistry Study Guide Answers
 - Interactive Elements Hydrocarbons Chemistry Study Guide Answers
8. Staying Engaged with Hydrocarbons Chemistry Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hydrocarbons Chemistry Study Guide Answers
9. Balancing eBooks and Physical Books Hydrocarbons Chemistry Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hydrocarbons Chemistry Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Hydrocarbons Chemistry Study Guide Answers
 - Setting Reading Goals Hydrocarbons Chemistry Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Hydrocarbons Chemistry Study Guide Answers
 - Fact-Checking eBook Content of Hydrocarbons Chemistry Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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