

Stoichiometry & Percent Yield Practice

- 1) Calculate the molar mass of each compound. Round to 3 decimal places.

a) NaCl 58.443 g/mol

b) FeCl_3 162.204 g/mol

c) $\text{Mg}(\text{NO}_3)_2$ 246.055 g/mol

- 2) $2 \text{AgNO}_3 + \text{Ni} \rightarrow 2 \text{Ag} + \text{Ni}(\text{NO}_3)_2$ BOTH SIGNIFICANTLY

- a) If 6.50 moles of nickel (Ni) reacts with excess silver nitrate (AgNO_3), how many moles of silver (Ag) will be produced?

$$6.50 \text{ mol Ni} \times \frac{2 \text{ mol Ag}}{1 \text{ mol Ni}} = \boxed{13.0 \text{ mol Ag}}$$

mole ratio

- b) If 2.35 moles of silver nitrate (AgNO_3) reacts completely, how many grams of silver (Ag) will be created?

$$2.35 \text{ mol AgNO}_3 \times \frac{2 \text{ mol Ag}}{2 \text{ mol AgNO}_3} \times \frac{107.868 \text{ g Ag}}{1 \text{ mol Ag}} = \boxed{253 \text{ g Ag}}$$

mole ratio

- c) If 4.13 g of Ni react with excess silver nitrate, how many moles of $\text{Ni}(\text{NO}_3)_2$ will be produced?

$$4.13 \text{ g Ni} \times \frac{1 \text{ mol Ni}}{58.693 \text{ g Ni}} \times \frac{1 \text{ mol Ni}(\text{NO}_3)_2}{1 \text{ mol Ni}} = \boxed{0.07 \text{ mol Ni}(\text{NO}_3)_2}$$

mole ratio

- d) If 100 g of Ni is produced from the reaction above, what mass of $\text{Ni}(\text{NO}_3)_2$ will also be produced?

$$100 \text{ g Ni} \times \frac{1 \text{ mol Ni}}{58.693 \text{ g Ni}} \times \frac{1 \text{ mol Ni}(\text{NO}_3)_2}{1 \text{ mol Ni}} \times \frac{246.055 \text{ g Ni}(\text{NO}_3)_2}{1 \text{ mol Ni}(\text{NO}_3)_2} = \boxed{1690 \text{ g Ni}(\text{NO}_3)_2}$$

mole ratio

Holt Stoichiometry Test Answers

A Loxley



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