Same Class Date

Assessment

Chapter Test B

Atomic Physics

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

What is the frequency of a photon with an energy of 1.99 × 10⁻¹⁹ J?

$$(h = 6.63 \times 10^{-14} \text{ Jes})$$

a.
$$1.00 \times 10^{14} \, \mathrm{Hz}$$

d.
$$4.00 \times 10^{14} \, \mathrm{Hz}$$

2. Light with an energy equal to three times the work function of a given metal causes the metal to eject photoelectrons. What is the ratio of the maximum photoelectron kinetic energy to the work function?

3. A monochromatic light beam with a quantum energy value of 3.0 eV is incident upon a photocell. The work function of the photocell is 1.6 eV. What is the maximum kinetic energy of the ejected electrons?

- 4. What causes the bright lines in the emission spectrum of an element to occur?
 - a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
 - b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
 - c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
 - d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.
 - S. What causes the dark lines in the absorption spectrum of an element to occur?
 - a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
 - b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
 - c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
 - d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

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