|  | lon  |  |   |   |   |
|--|--|--|---|---|---|
|  | each sentence or   | r statement.   |   |   |   |
| 12.  | An electron that of  | occupies the outer electron.   | most energy l                           | evel of an ato  | m is known as a(n)  |
| 13.  | An atom or group<br>one or more elect  |  | an electrical                           | charge becau  | se it has either lost or gained   |
| 14.  | An ion that has a  |  | sa(n) Chr                               | 0-  | 70 C  |
|  | The ion formed by  |  |   | ation   |   |
| 16.  | The ion formed by  | y an atom of a nor   | nmetal elemen                           | it is a(n)  | n on  |
|  | a(n)5 \( \sim / f \)   |  |   |   | veen a cation and an anion is   |
| 18.  | A repetitive geom  | etric arrangement  | t of ions, atom                         | s, or molecule  | es that forms a crystal structur  |
| 19.  | is called the _<   | ion $O^2$ is the   | xule.                                   | ion.  |   |
| 20.  | The name of the  | ion Cu* is the   | 0017-(1)                                | ion.  |   |
| 21.  | The chemical form  | mula for the comp  | ound strontiun                          |   | th contains Sr2+ and S2- ions,  |
|  | 5-5  |  |   |   |   |
| ord Ba   |  |  |   |   |   |
|  | Valence  | Crystal lattice<br>Salt  | Oxide                                   | Anion   | Copper(I)   |
|  |  |  |   |   |   |
| rite the   | formula for the f  | following tonic  | Write th                                | e names for   |   |
| rite the   |  | following lonic  | Write th                                |   | the following ionic   |
| mpou   |  | Nac1   | compou                                  |   |   |
| mpour<br>1. So   | nds:   | Nac1<br>Cao  | compou                                  | ınds:   | the following ionic   |
| 1. So<br>2. Ca<br>3. Iro                                   | nds:<br>dium chloride:<br>dcium oxide:<br>n (II) fluoride:   | No.CI<br>CaO<br>Fe F2  | сотроі<br>1<br>                         | ınds:   | the following ionic   |
| 1. So<br>2. Ca<br>3. Iro<br>4. Co                          | nds:<br>dium chloride:<br>dcium oxide:<br>n (II) fluoride:<br>apper (I) oxide:   | Nocl<br>Cao<br>Fe Fz<br>Cuzo   | сотроі<br>1<br>2                        | inds:<br>. NaNO <sub>3</sub>  | sodium Nitrate  |
| 1. So<br>2. Ga<br>3. Iro<br>4. Co<br>5. Alu                | nds:<br>dium chloride:<br>dcium oxide:<br>n (II) fluoride:   | No.CI<br>CaO<br>Fe F2  | 2 (2) (3) (3)                           | inds:<br>. NaNO3<br>. Li <sub>2</sub> O   | salium Nitrate  |
| 1. So<br>2. Ca<br>3. Iro<br>4. Co<br>5. Alu                | nds:<br>dium chloride:<br>dcium oxide:<br>n (II) fluoride:<br>apper (I) oxide:<br>uminum sulfate:                                    | No.CI<br>CaO<br>Fe F2<br>Cu20<br>A12 (50<br>A5 CI                          | 2 ( ) ( ) ( ) ( ) ( ) ( ) ( ) ( ) ( ) ( | . NaNO <sub>3</sub><br>. Li <sub>2</sub> O<br>. CuCO <sub>3</sub>   | lithium oxile  Copper-CII) carbon  Iron (III) oxide   |
| 1. So<br>2. Ca<br>3. Iro<br>4. Co<br>5. Alu<br>6. Sil      | nds: dium chloride: dcium oxide: n (II) fluoride: pper (I) oxide: uminum sulfate: ver chloride:                                      | No.Cl<br>Ca O<br>Fe F2<br>Cu2 O<br>Alz (SO<br>Ag Cl<br>Kz CO;<br>Sn Fy     | 2<br>2<br>2<br>3<br>4<br>5<br>6         | . NaNO <sub>3</sub> . Li <sub>2</sub> O . CuCO <sub>3</sub> . Fe <sub>2</sub> O <sub>3</sub>  | Sodium Nitrate  Lithium oxile  Copper-CID carbon  Iron (NI) oxide   |
| 1. So 2. Ca 3. Iro 4. Co 5. Alu 6. Silt 7. Po 8. Tir       | nds: dium chloride: dcium oxide: n (II) fluoride: pper (I) oxide: uminum sulfate: ver chloride: dassium carbonate n (IV) fluoride:   | NoCl<br>CaO<br>Fe Fz<br>Cuzo<br>Alz (SO<br>Ascl<br>Kz Coz<br>Sn Fy<br>Crzo | compoi<br>1<br>2<br>3<br>4<br>5         | inds: . NaNO <sub>3</sub> . Li <sub>2</sub> O . CuCO <sub>3</sub> . Fe <sub>2</sub> O <sub>3</sub> . Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>                | the following lonic  Salium Nitrate  Lithium oxile  Copper-CII) carboni  Iron (III) oxide  Calcium phosphate                                    |
| 1. So 2. Ca 3. Iro 4. Co 5. Alu 6. Silt 7. Po 8. Tir 9. Ch | nds: dium chloride: dicium oxide: n (II) fluoride: apper (I) oxide: uminum sulfate: ver chloride: dassium carbonate n (IV) fluoride: | NoCl<br>CaO<br>Fe Fz<br>Cuzo<br>Al, Cso<br>Ascl<br>KzCo,<br>Sm Fy          | compoi<br>1<br>2<br>3<br>5<br>6         | inds:  NaNO <sub>3</sub> Li <sub>2</sub> O  CuCO <sub>3</sub> Fe <sub>2</sub> O <sub>3</sub> Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> SnO                    | the following ionic  Salium Nitrate  Lithium oxile  Copper-CII) carbon  Iron (III) oxide  Calcium phosplate  tin (III) oxide                    |
| 1. So 2. Ca 3. Iro 4. Co 5. Alu 6. Silt 7. Po 8. Tir 9. Ch | nds: dium chloride: dcium oxide: n (II) fluoride: pper (I) oxide: uminum sulfate: ver chloride: dassium carbonate n (IV) fluoride:   | NoCl<br>CaO<br>Fe Fz<br>Cuzo<br>Alz (SO<br>Ascl<br>Kz Coz<br>Sn Fy<br>Crzo | compoi<br>1<br>2<br>3<br>5<br>6<br>6    | inds:  NaNO <sub>3</sub> Li <sub>2</sub> O  CuCO <sub>3</sub> Fe <sub>2</sub> O <sub>3</sub> Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> SnO  MgCO <sub>3</sub> | the following ionic  Sodium Nitrate  Lithium oxile  Copper (11) carbon  Iron (111) oxide  Calcium phosplate  tin (11) oxide  magnesium carbonat |

# Holt Science Spectrum Physical Science Answer Key Chemical Reactions

**JA Banks** 

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