

34. The maximum number of covalent bonds formed by nitrogen is

A. 1.
B. 2.
C. 3.
D. 4.

35. The IUPAC name of the compound $\text{CH}_3\text{CH}(\text{CH}_3)\text{CHCH}_2$ is

A. 2 - methyl but - 1 - ene.
B. 2 - methyl but - 2 - ene.
C. 3 - methyl but - 1 - ene.
D. 3 - methyl but - 2 - ene.

36. Ionization energy increases across the period in the periodic table because

A. atomic number increases.
B. effective nuclear charge increases.
C. mass number decreases.
D. screening effect decreases.

37. Which of the following properties indicate that an element is a metal. It

I. reacts with oxygen to form an acidic oxide
II. forms ionic chlorides.
III. has variable oxidation states.
IV. displaces hydrogen from dilute HCl.

A. I and III only
B. I and II only
C. II and IV only
D. I, II, III and IV

38. The electron configuration of carbon atom in its excited state is

A. $1s^2 2s^2 2p_x^1 2p_y^0$
B. $1s^2 2s^2 2p_x^1 2p_y^1$
C. $1s^2 2s^1 2p_x^1 2p_y^1 2p_z^1$
D. $1s^1 2s^2 2p_x^1 2p_y^1 2p_z^1$

39. An oxide has the following properties. It

I. is a white powder.
II. reacts with HCl.
III. reacts with NaOH.
IV. is insoluble in water.

The oxide is

A. alkaline.
B. amphoteric.
C. acidic.
D. neutral.

40. Which of the following statements about atoms of a metal is correct? They

A. readily accept electrons.
B. are soft.
C. are held together by covalent bond.
D. are held together by a sea of electron cloud.

41. Which of the following sketches is a graphical representation of Boyle's law?



42. The atom with the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^4$ is in

A. period 4, p- block.
B. period 3, p- block.
C. period 4, d- block.
D. period 3, d- block.

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Disha Experts

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Disha Experts' text.

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 - The Rise of Digital Reading May 2013 Chemistry C
 - Advantages of eBooks Over Traditional Books
2. Identifying May 2013 Chemistry C
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an May 2013 Chemistry C
 - User-Friendly Interface
4. Exploring eBook Recommendations from May 2013 Chemistry C
 - Personalized Recommendations
 - May 2013 Chemistry C User Reviews and Ratings
 - May 2013 Chemistry C and Bestseller Lists
5. Accessing May 2013 Chemistry C Free and Paid eBooks
 - May 2013 Chemistry C Public Domain eBooks
 - May 2013 Chemistry C eBook Subscription Services
 - May 2013 Chemistry C Budget-Friendly Options
6. Navigating May 2013 Chemistry C eBook Formats

- ePub, PDF, MOBI, and More
- May 2013 Chemistry C Compatibility with Devices
- May 2013 Chemistry C Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of May 2013 Chemistry C
 - Highlighting and Note-Taking May 2013 Chemistry C
 - Interactive Elements May 2013 Chemistry C
- 8. Staying Engaged with May 2013 Chemistry C
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers May 2013 Chemistry C
- 9. Balancing eBooks and Physical Books May 2013 Chemistry C
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection May 2013 Chemistry C
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine May 2013 Chemistry C
 - Setting Reading Goals May 2013 Chemistry C
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of May 2013 Chemistry C
 - Fact-Checking eBook Content of May 2013 Chemistry C
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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