

Section A

Answer all questions in the spaces provided

0 1

Anhydrous magnesium chloride, MgCl_2 , can absorb water to form the hydrated salt $\text{MgCl}_2 \cdot 4\text{H}_2\text{O}$



← creating crystals directly is not really possible.

0 1 - 1

Suggest one reason why the enthalpy change for this reaction cannot be determined directly by calorimetry.

[1 mark]

Virtually impossible to prevent dissolving.

0 1 - 2

Some enthalpies of solution are shown in Table 1.

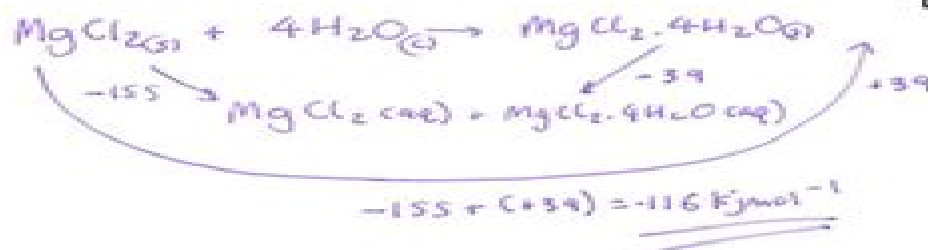
Table 1

Salt	Enthalpy of solution / kJ mol^{-1}
$\text{MgCl}_2(\text{s})$	-155
$\text{MgCl}_2 \cdot 4\text{H}_2\text{O}(\text{s})$	-39

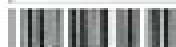
making the salts aqueous!

Calculate the enthalpy change for the absorption of water by $\text{MgCl}_2(\text{s})$ to form $\text{MgCl}_2 \cdot 4\text{H}_2\text{O}(\text{s})$.

[2 marks]



Enthalpy change _____ kJ mol^{-1}



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Guillaume Favre

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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