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Jiyuan Zhang



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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example $\text{hcl naoh nacl h}_2\text{o}$

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web during this reaction the h^+ ions from the acid and the oh^- ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable
measure an appropriate volume of each liquid eg 25 cm³ place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less
exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of
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eg 1 hcl aq + naoh aq → nacl aq + h₂O l for which ΔH = -57 kJ mol⁻¹ eg 2 hNO₃ aq + KOH aq → KNO₃ aq + H₂O l for which ΔH = -57 kJ mol⁻¹

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react
together to produce one mole of water for example hcl aq + NaOH aq → NaCl aq + H₂O aq you can see that one mole of water is
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