UNIT REVIEW - BONDING

Intramolecular Forces

- 1. Define electronegativity
 - the tendency of atoms to attract e-
- as you go across the periodic table, and <u>clearearer</u> as you go 2. Electronegativity (ncreous) down a group.
- 3. Explain why atoms with smaller atomic radii have stronger electronegativities

 - smaller atoms means e's are closer to the nucleus, which has the protons if you decrease the distance between protons + e's, you increase attractive forces.
- 4. Define intramolecular force
 - forces that hold atoms together ionic, covalent, polar covalent bond
- Compare and contrast covalent, polar covalent, and ionic bond.

Covalent	Polar covalent	lonic
- e- are shared equally - between a non-motall - difference in electronegativity is = 0.2 - atoms are newbal	- e - are shared unequally - between a non-metals - difference in electronegotisty is 0.3-1.6 - atoms are partially charged	- e are transfered - between metal + nonmeta - difference in electronegativity is ≥1.7 - atoms have a charge

6. What is a partial charge and how is it different from an ionic charge?

A partial charge results from an unequal shaving of elections between a atoms An ionic change results from a transfer of elections from I atom to another

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D Keegan

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web wed may 1st 2024 school principals day organize an event or write a card to show your appreciation for the educators that run our schools school principals work hard to make kids lives and educations better have we missed something you can have great teachers but if you don't have a good principal you won't have a good school eli broad

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