

## Models of Covalent Bonding

Molecules have shape! The structure and shape of a molecule influences its physical properties and affects its chemical behavior as well. Lewis structures and VSEPR theory offer useful models for visualizing the structure of covalent compounds.

- Valence electrons
- Lewis structures
- Covalent bonding
- VSEPR theory

Covalent bonds are defined as the net attractive forces between electrons of all atoms that share one, two or three pairs of electrons. In general, only the valence electrons, those in the highest energy levels that are farthest away from the nucleus, are available for bonding. The number of valence electrons influences the number of bonds that an atom will form. The Periodic Table offers a convenient shortcut for determining the number of valence electrons in an atom. Remember that the position of an element in the modern Periodic Table reflects its electron configuration. When the representative elements are arranged in columns from IA to VIIA (see Figure 1), the number of valence electrons for an element is equal to its group number. Thus, potassium in Group IA has one valence electron, carbon in Group IVA has four valence electrons, and chlorine in Group VIIA has seven valence electrons.

Figure 1. Nomenclature of stereoregulating elements in Cations Ia-VIIa

In 1916, G.N. Lewis, an American chemist, proposed arranging dots around the symbols of the elements to represent the valence electrons. *Lewis electron dot symbols* (see Figure 2) remain the most popular way to picture the valence electrons that are available for bonding.

Figure 2. Lysis characteristics as related to concentration of elements.

Lewis structures based on the Lewis structures (or symbols) of the elements to show the bonding arrangement of atoms in a molecule and the distribution of all valence electrons. The Lewis structure of a molecule thus shows all of the atoms and how they are connected. A single covalent bond between two atoms, corresponding to a pair of electrons, is represented using a dash (—). Sometimes atoms have more than one pair of electrons between them in order to form stable molecules. Two dashes (corresponding to two pairs of electrons) and three dashes (corresponding to three pairs of electrons) are used to represent double and triple bonds, respectively.

G. N. Fano's offered a simple theory, based on the known stability of the noble gases (Fano, 1962), namely, the prediction that many bonded atoms will form and lose many atoms of a particular type still come together to form a stable molecule. According to Fano's, nonmetallic many-atom electronegative anion to achieve a valence shell electron "count," similar to that of the noble gases:

"2. Use of a name may conform to the rule of right, or the water rule, by changing over or more pairs of abstracts. 3. The abstracts which are listed in common by two systems may be considered to belong to the water shell of both systems."

The noble gases have filled  $s$  and  $p$  orbitals with eight electrons. The octet rule assumes that atoms form molecules to achieve this stable, noble gas electron configuration. In counting valence electrons in order to predict the structure of a covalent compound, we will distinguish between two kinds of electron pairs. *Bonding pairs* of electrons are shared between atoms and thus "belong" to both atoms in the bond. *Nonbonding or lone pairs* of electrons are not shared between atoms and are therefore "owned" to one or both of the atoms.

# Flinn Chem Topics Lewis Structures And Molecular Geometry

**F Rizvi**

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