

100) Carbonic acid reacts as a buffer at pH 7.4. How many moles of H_2CO_3 react with 1 mole of HCO_3^- to form 1 mole of H_2O and 1 mole of CO_2 ?



Answer: $1 \text{ mole H}_2\text{CO}_3$ reacts with 1 mole HCO_3^- to form $1 \text{ mole H}_2\text{O}$ and 1 mole CO_2 .

101) The only intracellular buffer in body fluid is not the phosphate.

102) Describe how the phosphate buffer works.

Phosphate buffer works by reacting with H^+ ions to form H_2PO_4^- and HPO_4^{2-} ions. The HPO_4^{2-} ion is the conjugate base of H_2PO_4^- and the H_2PO_4^- ion is the conjugate acid of HPO_4^{2-} .

103) Describe how the phosphate buffer works in body fluid.

The phosphate buffer works by reacting with H^+ ions to form H_2PO_4^- and HPO_4^{2-} ions. The HPO_4^{2-} ion is the conjugate base of H_2PO_4^- and the H_2PO_4^- ion is the conjugate acid of HPO_4^{2-} .

104) The phosphate buffer is a component of body fluid. Describe its function.

Describe how the phosphate buffer works in body fluid.

What is the conjugate base of H_2PO_4^- ?

Answer: HPO_4^{2-}



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