change of the surrounding water. $\Delta H_{metal} = -\Delta H_{water}$ If the water temperature increases, then the metal must have released heat. We can find the enthalpy change of the metal with the formula: $\Delta H_{metal} = -m_{water} \cdot C_{water} \cdot \Delta T_{water}$

Check-for-understanding question:

When 25 mL of water containing 0.85 mol HCl is mixed with 40 mL of water containing .05 mol KOH in a calorimeter (with initial temperature 23 °C), the temperature of the solution increases to 25.5 °C. How many kilojoules (kJ) of heat are released by the acid-base reaction?

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