



INSTALLER'S SYSTEM SETUP GUIDE

icomfort™ Thermostat

Touch Screen Programmable Communicating Thermostat

CONTROLS
506052-01
11/09

TTP Technical
Publications
Litho U.S.A.

Shipping and Packing List1

- 1 - icomfort™ Touch Screen Communicating, 7-day Programmable Thermostat
- 6 - Mounting Screws
- 6 - Wall Anchors
- 1 each - Installation Quick-Start Guide, Installer's System Setup Guide, Homeowner's Manual, Warranty card, Warranty Audit tag

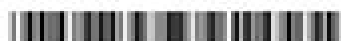


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WARNING

Always turn off power at the main power source by switching the circuit breaker to the OFF position before installing or removing this thermostat.

All wiring must conform to local and national building and electrical codes and ordinances.

Do not switch system to cool if the outdoor temperature is below 45°F (7°C). This can damage the cooling system.



CAUTION

This is a 24VAC low-voltage thermostat. Do not install on voltages higher than 30VAC.

Do not short (jumper) across terminals on the gas valve or at the system control to test installation. This will damage the thermostat and void the warranty.



IMPORTANT

Read this manual before programming this thermostat.
Use this thermostat only as described in this manual.

Lennox Thermostat Installation Manual

JL Elias



Lennox Thermostat Installation Manual:

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a usia kehamilan 1 klasifikasi menurut usia kehamilan bayi diklasifikasikan menjadi bayi kurang bulan bayi cukup bulan dan bayi lebih bulan bayi kurang bulan bkb adalah bayi

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dalam generasi yang solid dikatakan bahwa usia baik baik saja untuk kehamilan dan persalinan pada usia 20 tahun hingga 30

tahun meskipun kematian dalam persalinan dikatakan berada

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fertilisasi sampai kelahiran bayi kehamilan normal biasanya berlangsung dalam waktu 40 minggu usia kehamilan tersebut dibagi menjadi 3 trimester yang masing masing

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hidup di luar kandungan dengan berat badan kurang dari 1000 gram atau usia kehamilan kurang dari 28 minggu manuba 2010 dapat disimpulkan bahwa abortus adalah

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1 bayi prematur preterm kurang bulan usia kehamilan kurang dari 37 minggu 259 hari 2 bayi cukup bulan aterm usia kehamilan antara 37 dan 42 minggu 259 293 hari 3 bayi

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hari 40 minggu dan tidak lebih dari 300 hari 43 minggu pembagian kehamilan dibagi dalam 3 trimester trimester i dimulai dari konsepsi sampai tiga bulan 0 12 minggu trimester ii

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web 1 consider the reaction $\text{I}_2\text{O}_5 + 5 \text{CO} \rightarrow 5 \text{CO}_2 + \text{I}_2$ a 80 0 grams of iodine v oxide I_2O_5 reacts with 28 0 grams of carbon

monoxide co determine the mass of iodine i₂ which could be produced b if in the above situation only 0.160 moles of iodine i₂ was produced i what mass of iodine was produced

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole h₂s and multiply it with 2 mole h₂o over 2 mole h₂s from the equation

$$8.3 \text{ mole h}_2\text{s} \times \frac{2 \text{ mole h}_2\text{o}}{2 \text{ mole h}_2\text{s}} = 8.3 \text{ mole h}_2\text{o}$$

 8.3 mole h₂s x 18g h₂o i mole h₂o 149.4g h₂o is your theoretical yield so percent yield 137.1g h₂o actual yield 149.4g h₂o x 100
 91.77 problems 6 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21.3 grams of o₂ produces a 47.3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent

you have 0.216232 grams mole 21.3 grams 0.66 moles 32 g m

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0

g $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol $\text{Al} \times 1 \text{ mol Al}_2\text{O}_3$ c determine the percent yield

yield actual $\times 100$ $260.0 \text{ g} \times 100 / 86.01 \text{ theoretical} = 302.3 \text{ g}$

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

C is the limiting reagent 4 33 g of H₂ are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s

remains when the reaction is complete

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent

yield theoretical $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ based on this definition we would expect a percent yield

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CH₃OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH₃OH if 68.5 g of CO is

reacted with 8.6 g of H₂

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problems from a balanced chemical equation 3 sto 5 differentiate between the actual yield and theoretical yield of a chemical

reaction

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