

(iii) The amount of acid added to the tablet in **Step 1** was 2.00×10^{-2} mol. Use this information and your answer to (b)(ii) to calculate the number of moles of hydrochloric acid that reacted with the tablet.

(1)

(iv) The equation for the reaction of the magnesium hydroxide in the tablet with hydrochloric acid is shown below.



Calculate the mass of magnesium hydroxide in the tablet. Give your answer to **three** significant figures.

The molar mass of magnesium hydroxide is 58.3 g mol^{-1} .

(2)

(c) The volume of hydrochloric acid added to the tablet was 40.00 cm^3 .

(i) Suggest a change in the procedure which would make the result of the experiment **more reliable** for each tablet which is analysed.

(1)



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