

Example 1

Calculate the K_a value of a 0.03 M aqueous solution of hydrofluoric acid (HF) with a pH of 2.75.



	HF	H ₂ O	→	F ⁻	H ₃ O ⁺
I	0.03	~		0	0
C	-x			+x	+x
E	0.03-x			+x	+x

pH = 2.75

$$[\text{H}_3\text{O}^{+}] = 10^{-2.75} = 1.8 \times 10^{-3} \text{ M}$$

$$[\text{F}^{-}] = 1.8 \times 10^{-3} \text{ M}$$

$$K_a = \frac{[\text{products}]}{[\text{reactants}]} = \frac{[\text{F}^{-}][\text{H}_3\text{O}^{+}]}{[\text{HF}]}$$

$$K_a = \frac{(x)(x)}{(0.03-x)} = \frac{x^2}{(0.03-x)}$$

$$K_a = \frac{(1.8 \times 10^{-3})^2}{(0.03 - (1.8 \times 10^{-3}))} = 1 \times 10^{-4}$$

Ka Practice Chemistry

G Thomas



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